

## Example: Dimensional Analysis – Stoichiometry and mass-to-mass conversions

How many grams of carbon monoxide are produced from 46.5 g FeTiO<sub>3</sub> (ilmenite) according to the equation:



$$46.5 \text{ g } \cancel{\text{FeTiO}_3} \times \frac{1 \text{ mol } \cancel{\text{FeTiO}_3}}{151.73 \text{ g } \cancel{\text{FeTiO}_3}} \times \frac{3 \text{ mol } \cancel{\text{CO}}}{1 \text{ mol } \cancel{\text{FeTiO}_3}} \times \frac{28.01 \text{ g } \text{CO}}{1 \text{ mol } \cancel{\text{CO}}} \\ = 25.8 \text{ g CO}$$

# Chem 101 Activity On Dimensional Analysis Answers

**SA Adler**



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