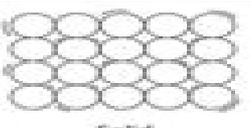


# 1. The Particulate Nature of Matter

## 1.1. Kinetic Particle Theory

		
<b>Solid</b>	<b>Liquid</b>	<b>Gas</b>
Strong intermolecular forces	Weaker intermolecular forces than solids	Almost no intermolecular forces
Fixed lattice arrangement	No fixed arrangement; particles can move and slide over each other	Particles far apart and move quickly
Particles vibrate in fixed position; fixed shape and volume	Particles slide; fixed volume	Random movement; no fixed shape or volume

- When a solid is heated, the particles gain sufficient energy to overcome the strong intermolecular forces. The particles eventually can slide over each other in a more random motion- solid expands until the structure is broken at m.p.
- When a liquid is heated to its b.p, the particles overcome the relatively weaker intermolecular force to escape the liquid's surface and move around in continuous rapid motion – the liquid has boiled
- In the vapor, the particles move in rapid random motion. This movement is due to the collision of vapor particles with air particles.
- When gaseous particles are heated in a closed environment, the increase in kinetic energy causes increased collisions with other particles as well as the walls of the container- pressure increases.

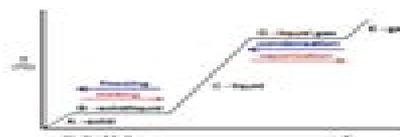
## 1.2. States of Matter



Process	Change	Heat Energy	Exo/endothermic
Melting	S → L	Gained	Endothermic
Boiling	L → G	Gained	Endothermic
Condensing	G → L	Lost	Exothermic

Process	Change	Heat Energy	Exo/endothermic
Freezing	L → S	Lost	Exothermic
Sublimation	S → G	Gained	Endothermic
Reverse sublimation	G → S	Lost	Exothermic

## 1.3. Heating Curve



## 1.4. Brownian motion and Diffusion

### 1. Brownian motion:

- It is the random movement of particles in a liquid or a gas caused due to collision with smaller, invisible particles
- Evidence: In liquid- Pollen grains in water  
In gases- Smoke in air

### 1. Diffusion

- It is the spreading of one substance (liquid or gas) through another from a region of high concentration to a region of low concentration due to the continuous random motion of particles.
- Evidence for diffusion:
  - In liquids: potassium manganate (VII) in a beaker of water
  - In gases: a gas jar of air and a gas jar of bromine connected
- Factors that affect the rate of diffusion:
  - Temperature increases → rate of diffusion increases
  - Lower relative molecular mass → rate of diffusion is higher



# 2. Experimental Techniques

## 2.1. Measurement

Variable	Unit	Apparatus
Time	min/sec	Stopwatch
Temperature	°C	Thermometer (liquid in glass, thermistor or thermocouple)
Mass	grams	Balance

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