

## Section 12.1

### The Arithmetic of Equations

#### OBJECTIVES:

- Explain how balanced equations apply to both chemistry and everyday life.
- Interpret balanced chemical equations in terms of: a) moles, b) representative particles, c) mass, and d) gas volume (Liters) at STP.
- Identify the quantities that are always conserved in chemical reactions.

#### Stoichiometry is...

- Greek for "measuring elements"
- Pronounced "stoy keg ahm uh tree"
- Defined as: \_\_\_\_\_
- 

- There are 4 ways to *interpret* a balanced chemical equation

#### #1. In terms of Particles

- An Element is made of \_\_\_\_\_
- A Molecular compound (made of only nonmetals) is made up of \_\_\_\_\_  
(Don't forget the diatomic elements)
- Ionic Compounds (made of a metal and nonmetal parts) are made of \_\_\_\_\_

Example:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

- Two *molecules* of hydrogen and one *molecule* of oxygen form two *molecules* of water.
- Another example:  $2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$

Now read this:  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$

#### #2. In terms of Moles

- The coefficients tell us how many moles of each substance
- $$2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$$

# Chemistry Chapter 12 Stoichiometry Notes

**Raffaella Di Napoli**



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