

Chemistry 5A

Experiment 3

Introduction to Density

Introduction

The purpose of this experiment is to understand the meaning and significance of the density of a substance. Density is a basic physical property of a homogeneous substance; it is an intensive property, which means it depends only on the substance's composition and does not vary with size or amount. The determination of density is a nondestructive physical process for distinguishing one substance from another. Density is the ratio of a substance's mass to its total volume:

$$d = \frac{m}{V} = \frac{\text{mass}}{\text{volume}}$$

In the metric system the unit of density has a liquid or solid as measured in grams, or grams³. The unit volume used with solids is conventionally equal to mL, volume unit used with liquids. That is, 1 mL = 1 cm³. In this experiment you will determine the density of several liquids and compare the physical properties of these liquids.

Which is heavier, a pound of aluminum or a pound of lead? The answer, of course, is neither, but many people consider the words "heavy" and "dense" "heavy" refers to mass only. Density is the mass of a substance contained in a unit of volume. Lead is a very dense metal and contains a large quantity of matter in a small volume, while aluminum, being much less dense, contains a smaller quantity of matter in the same volume.

The volume of 200 grams of lead is 1.77 mL. The mass of lead contained in each mL is its density.

$$\text{Density of lead} = \frac{200\text{g}}{1.77\text{ mL}} = \frac{200\text{g}}{1.77\text{ mL}} = 113\text{g/cm}^3$$

The volume of 200 g of aluminum is 7.41 mL.

$$\text{Density of aluminum} = \frac{200\text{g}}{7.41\text{ mL}} = 27\text{g/cm}^3$$

From the definition of the gram and the milliliter, we can say that one mL of water at 4 °C would have a mass of exactly one gram. The density of water, then, is one gram, or 1 g/cm³. Since the volume occupied by one gram of water varies slightly with temperature, the density also varies slightly with changes in temperature.

The mass of any object is determined by comparing its mass with the mass of known object or objects (i.e., it is weighed). The volume of a liquid is measured using a graduated cylinder, a pipet, or some other

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Peter P. Berlow



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Table of Contents Chemistry 51 Experiment 3 Introduction To Density

1. Understanding the eBook Chemistry 51 Experiment 3 Introduction To Density
 - The Rise of Digital Reading Chemistry 51 Experiment 3 Introduction To Density
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 51 Experiment 3 Introduction To Density
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 51 Experiment 3 Introduction To Density
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 51 Experiment 3 Introduction To Density
 - Personalized Recommendations
 - Chemistry 51 Experiment 3 Introduction To Density User Reviews and Ratings
 - Chemistry 51 Experiment 3 Introduction To Density and Bestseller Lists
5. Accessing Chemistry 51 Experiment 3 Introduction To Density Free and Paid eBooks
 - Chemistry 51 Experiment 3 Introduction To Density Public Domain eBooks
 - Chemistry 51 Experiment 3 Introduction To Density eBook Subscription Services
 - Chemistry 51 Experiment 3 Introduction To Density Budget-Friendly Options

6. Navigating Chemistry 51 Experiment 3 Introduction To Density eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry 51 Experiment 3 Introduction To Density Compatibility with Devices
 - Chemistry 51 Experiment 3 Introduction To Density Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 51 Experiment 3 Introduction To Density
 - Highlighting and Note-Taking Chemistry 51 Experiment 3 Introduction To Density
 - Interactive Elements Chemistry 51 Experiment 3 Introduction To Density
8. Staying Engaged with Chemistry 51 Experiment 3 Introduction To Density
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 51 Experiment 3 Introduction To Density
9. Balancing eBooks and Physical Books Chemistry 51 Experiment 3 Introduction To Density
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 51 Experiment 3 Introduction To Density
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 51 Experiment 3 Introduction To Density
 - Setting Reading Goals Chemistry 51 Experiment 3 Introduction To Density
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 51 Experiment 3 Introduction To Density
 - Fact-Checking eBook Content of Chemistry 51 Experiment 3 Introduction To Density
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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- Interactive and Gamified eBooks

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