

## Answers to Chapter 3 Study Questions

1. a)  $3(12.0) + 8(1.01) + 3(16.0) = 92.1 \text{ g/mole}$       b)  $92.1 \text{ g}$       c)  $6.02 \times 10^{23}$  molecules  
 d)  $0.217 \text{ moles} \times \frac{92.1 \text{ g}}{1 \text{ mole}} = 20.0 \text{ g}$
2. a) 4 atoms (one N + 3 H)  
 b)  $1 \text{ mole } \text{NH}_3 \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 2.41 \times 10^{24} \text{ atoms}$   
 c)  $3.40 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 4.82 \times 10^{23} \text{ atoms}$
3. Molar mass of  $\text{NaNO}_2 = 23.0 + 14.0 + 2(16.0) = 69.0 \text{ g/mole}$   
 $\% \text{ Na} = 23.0/69.0 = 33.3\% \text{ Na}$ ;  $\% \text{ N} = 14.0/69.0 = 20.3\% \text{ N}$ ;  $\% \text{ O} = 2(16.0)/69.0 = 46.4\% \text{ O}$   
**33.3% Na, 20.3% N and 46.4% O.**
4. a) In 100 g of this compound, there are 40.7 g C, 5.1 g H, and 54.2 g O  
 $40.7 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 3.39 \text{ moles C}$        $3.39/3.39 = 1 \times 2 = 2$   
 $5.1 \text{ g H} \times \frac{1 \text{ mole H}}{1.0 \text{ g H}} = 5.1 \text{ moles H}$        $5.1/3.39 = 1.5 \times 2 = 3$   
 $54.2 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 3.39 \text{ moles O}$        $3.39/3.39 = 1 \times 2 = 2$   
 empirical formula =  **$\text{C}_2\text{H}_3\text{O}_2$**
- b) Molar mass of  $\text{C}_2\text{H}_3\text{O}_2 = 2(12.0) + 3(1.0) + 2(16.0) = 59.0 \text{ g/mole}$   
 $118/59.0 = 2 \rightarrow$  molecular formula =  **$\text{C}_4\text{H}_6\text{O}_4$**
5. a) In 25.0 g of compound, there are 7.20 g Mg, 3.55 g C and  $25.0 - (7.20 + 3.55) = 14.25 \text{ g O}$ .  
 $7.20 \text{ g Mg} \times \frac{1 \text{ mole Mg}}{24.3 \text{ g Mg}} = 0.296 \text{ moles Mg}$        $0.296/0.296 = 1$   
 $3.55 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 0.296 \text{ moles C}$        $0.296/0.296 = 1$   
 $14.25 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 0.891 \text{ moles O}$        $0.891/0.296 = 3$   
 formula =  **$\text{MgCO}_3$**
- b)  $\% \text{ Mg} = 7.20/25.0 = 28.8\% \text{ Mg}$ ;  $\% \text{ C} = 3.55/25.0 = 14.2\% \text{ C}$ ;  $\% \text{ O} = 14.25/25.0 = 57.0\% \text{ O}$ . (You should get the same result using molar mass and atomic masses.)
- c)  $13.9 \text{ g compound} \times \frac{28.8 \text{ g Mg}}{100 \text{ g compound}} = 4.00 \text{ g Mg}$

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**Yijin Wang**



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## **Table of Contents Chemistry Chapter 3 Worksheet Answers**

1. Understanding the eBook Chemistry Chapter 3 Worksheet Answers
  - The Rise of Digital Reading Chemistry Chapter 3 Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 3 Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 3 Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 3 Worksheet Answers
  - Personalized Recommendations
  - Chemistry Chapter 3 Worksheet Answers User Reviews and Ratings
  - Chemistry Chapter 3 Worksheet Answers and Bestseller Lists
5. Accessing Chemistry Chapter 3 Worksheet Answers Free and Paid eBooks

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  - Chemistry Chapter 3 Worksheet Answers Budget-Friendly Options
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    - ePub, PDF, MOBI, and More
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    - Chemistry Chapter 3 Worksheet Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Chapter 3 Worksheet Answers
    - Highlighting and Note-Taking Chemistry Chapter 3 Worksheet Answers
    - Interactive Elements Chemistry Chapter 3 Worksheet Answers
  8. Staying Engaged with Chemistry Chapter 3 Worksheet Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Chapter 3 Worksheet Answers
  9. Balancing eBooks and Physical Books Chemistry Chapter 3 Worksheet Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Chapter 3 Worksheet Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Chapter 3 Worksheet Answers
    - Setting Reading Goals Chemistry Chapter 3 Worksheet Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Chapter 3 Worksheet Answers
    - Fact-Checking eBook Content of Chemistry Chapter 3 Worksheet Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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