

Name: \_\_\_\_\_

Date: \_\_\_\_\_

### Chapter 6 Study Guide: Periodicity

1. Describe the historical development of the periodic table (Newland, Meyer, Mendeleev, Mosely) Describe the contributions made by each of the scientist.

**Newland** – First to order the elements in a periodic table –developed the law of octaves

**Meyers** – Ordered elements by atomic mass but is given little credit for the current periodic table

**Mendeleev** – ordered elements by atomic mass and is given the most credit for the current periodic table because he left holes in his table for undiscovered elements and described their properties

**Mosely** – ordered elements by atomic number (modern periodic table)

2. Fill in the blanks below:

The elements in the periodic table are arranged in order of increasing **Atomic** \_\_\_\_\_ **Number** \_\_\_\_\_. The vertical columns are called **Groups** \_\_\_\_\_ or **Families** \_\_\_\_\_. The horizontal rows are called **Periods** \_\_\_\_\_. The elements within the same **Family** \_\_\_\_\_ have the same chemical properties due to the number of their **Valence** \_\_\_\_\_ **Electrons** \_\_\_\_\_. Metals are located on the **Left** \_\_\_\_\_ side of the periodic table. The nonmetals are located on the **Right** \_\_\_\_\_ side of the periodic table. The metalloids are located on the **Stair step** \_\_\_\_\_ between the metals and nonmetals.

3. Describe the general properties of the following groups of elements on the periodic table:

a. Alkali Metals

**1 valence electrons, metals, highly reactive, s-block**

b. Alkaline Earth metals

**2 valence electrons, metals, reactive, s-block**

c. Halogens

**7 valence electrons, high reactive, nonmetals, p-block**

d. Noble Gases

**8 noble electrons, nonreactive, nonmetals, all gases, p-block**

e. Transition Metals

**Variable valence electrons, metals, malleable, ductile, metals, conductive**

f. Inner Transition Metals

**Variable valence electrons, metals, f-block**

# Chemistry Chapter 6 Study Guide For Content Mastery Answers

**N Noddings**



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