

Conversion factors:

Molar mass = #g/1 mole

Avogadro's number = 6.02×10^{23} atoms or molecules/1 mole

Molarity = #moles/1L

Name _____

Period _____

Stoichiometry Practice Problem Set

- In the following reaction $4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s)$, iron oxidizes to form rust.
 - If 40.0g of iron react, how many moles of Fe_2O_3 , iron (III) oxide are produced?
 - If 20.0g of iron react, how many grams of Fe_2O_3 are produced?
 - If 10.0 moles of iron react, how many molecules of Fe_2O_3 are produced?
 - If 3.00×10^{22} atoms of iron react, how many moles of Fe_2O_3 are produced?
- In the reaction for photosynthesis, $6\text{CO}_2(g) + 6\text{H}_2\text{O}(l) + \text{light energy} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6(aq) + 6\text{O}_2(g)$,
 - If 715g of CO_2 gas is consumed, how many grams of glucose ($\text{C}_6\text{H}_{12}\text{O}_6$) are produced? The molar mass of glucose is 180.16g/mol.
 - How many molecules of carbon dioxide are required to produce 5.9×10^{24} molecules of oxygen gas?

Chemistry Unit 5 Stoichiometry Practice Problems I

State University of Iowa



Chemistry Unit 5 Stoichiometry Practice Problems I:

Resources in Education, 1980-04 *Reviewing Chemistry* Peter E. Demmin, 1989 **Solving Problems in Chemistry** Rod O'Connor, Charles Mickey, Alton Hassell, 1977 Essentials of Introductory Chemistry Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids and Bases For all readers interested in introductory chemistry *Introductory Chemistry* Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place tm Special Edition and on The Chemistry of Life CD ROM For college instructors and students **Resources in Education**, 1980 **Chemical & Metallurgical Engineering** Eugene Franz Roeber, Howard Coon Parmelee, 1926 Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **The Addison-Wesley Science Handbook** Gordon J. Coleman, David C. Dewar, 1997 Brings together a broad range of essential science information Both fundamental and advanced concepts are presented in table glossaries and summaries for quick memory refreshers at all levels **Electrochemical and Metallurgical Industry** Eugene Franz Roeber, Howard Coon Parmelee, 1926 **The Complete Idiot's Guide to Chemistry** Ian Guch, 2006 With topics such as stoichiometry chemical equilibria and thermodynamics it's no wonder so many students have a bad reaction to chemistry class Fortunately The Complete Idiot's Guide to Chemistry Second Edition gives readers an easy to understand formula for success when it comes to passing their chemistry class with flying colours Intended for both high school students who are required to take a college prep chemistry course and college students who need to take chemistry to meet their degree requirements this book gives readers a non intimidating easy to understand

companion to their textbooks The book follows a standard math based chemistry curriculum covering the following topics Atomic structure and elements The periodic table Ionic and covalent compounds Chemical reactions Kinetics Acids and bases Electrochemistry Organic chemistry Nuclear chemistry Stoichiometry Thermodynamics

Improving Student Comprehension in Chemistry Laboratories Tracy Lynn Haroff,2006

A Complete Preparation for the MCAT Aftab S. Hassan,James L. Flowers,1992 This guide for MCAT preparation applies the principles of active and problem based learning to an updated review of content and skills with models for enhanced problem solving and critical thinking abilities There are details on setting up a self managed study programme with guidelines for time management and stress management All areas tested on the exam are covered verbal reasoning physical science writing sample biological sciences with practice questions to chart progress

Civil Engineering Reference Manual Michael R. Lindeburg,1992 This manual fully prepares applicants for the civil PE exam by far the most popular of the PE disciplines Every exam subject is thoroughly covered with illustrations and practice problems to heighten the reader s understanding Also included are test taking strategies and exam information indexed

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