

To receive full credit, show all work with units for each question. Refer to the table below for appropriate constants, boiling points, etc.

Compound	Formula	Density (20 °C, g cm <sup>-3</sup> )	FP (°C)	BP (°C)	K <sub>f</sub> (°C/m)	K <sub>b</sub> (°C/m)
Acetic Acid	CH <sub>3</sub> COOH	1.0492	16.6	117.9	-3.90	3.07
Benzene	C <sub>6</sub> H <sub>6</sub>	0.8765	5.5	80.1	-4.90	2.53
Carbon Tetrachloride	CCl <sub>4</sub>	1.5940	-22.99	76.54	-2.98	5.03
Nitrobenzene	C <sub>6</sub> H <sub>5</sub> NO <sub>2</sub>	1.2037	5.7	210.8	-7.00	5.24
Water	H <sub>2</sub> O	0.998203	0.000	100.0	-1.86	0.512
Ethanol	C <sub>2</sub> H <sub>5</sub> O	0.7893	-117.3	78.5	-1.99	1.22

- For a mixture prepared by mixing 5.00 g of caffeine (C<sub>8</sub>H<sub>10</sub>N<sub>4</sub>O<sub>2</sub>) with 250.0 g of water
  - Calculate the boiling point of this mixture
  - Calculate the freezing point of this mixture
- For a mixture prepared by mixing 10.0 g of sodium sulfate in 250 g of water
  - Calculate the boiling point of this mixture
  - Calculate the freezing point of this mixture
- You are examining forensics data for a court case and need to identify an analgesic (pain killer). It could be either acetylsalicylic acid (HC<sub>9</sub>H<sub>7</sub>O<sub>4</sub>, 180.160 g/mol), ibuprofen (C<sub>13</sub>H<sub>18</sub>O<sub>2</sub>, 206.3 g/mol) or acetaminophen (C<sub>8</sub>H<sub>9</sub>NO<sub>2</sub>, 151.17 g/mol). A mixture of the unknown is prepared by mixing 10.0 grams of unknown and diluting to 250 mL with ethanol. The boiling point of this mixture is 78.8 °C. The density of this solution is 0.80 g/mL. Assume that the unknown takes up no volume in the solution. Determine the identity of the unknown. Big hint: Can you calculate the molar mass?

# Calculations Involving Colligative Properties Section Review Answers

**Gary Leighton Long, Sara Long, Robert  
C. Fay**

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