

Experiment 7

When performing acid-base titrations, the first step is often a standardization of the titrant, which is an accurate determination of its concentration. Acidic titrants might be standardized in titration of a well-defined amount of a solid base, such as sodium carbonate. In such a procedure, why is it not necessary to know the exact amount of water used in preparation of a solution of Na_2CO_3 used in standardization of an acid?

The standardization of an acid might also be carried out using a solution of a base with known concentration. For example, it is found that 37.60 ml of 0.210 M NaOH are required to neutralize 25.05 ml of H_2SO_4 solution in a titration experiment. Calculate the molarity and normality of the H_2SO_4 solution.

In the example given above, the amount of acid to be titrated has to be delivered by a pipet so that its volume is accurately known. However, the solution of the acid is often diluted to facilitate the titration process. Explain why additional de-ionized water might be added without affecting the results of the titration.

Experiment 7 Acid Base Titrations Answers

Mary Millhollon, Richard H. Langley



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