

Appendix D

ANSWERS TO SOME PROBLEMS

1-1. (a) At standard temperature and pressure, a mole of an ideal gas contains 6.022×10^{23} molecules (Avogadro's number) and occupies 22.4 liters. Hence, the number per m^3 is $6.022 \times 10^{23}/2.24 \times 10^{-2} = 2.66 \times 10^{25} \text{ m}^{-3}$.

(b) Since $PV = NRT$, $n = N/V = P/RT$. Hence $n_1/n_0 = P_1T_0/P_0T_1$. Taking n_0 to be the density in part (a) and n_1 to be that in part (b), we have

$$n_1 = (2.66 \times 10^{25}) \frac{10^{-3}}{760} \frac{273}{(273 + 20)} = 3.30 \times 10^{24} \text{ m}^{-3}$$

Note that a diatomic gas such as H_2 will have twice as many atoms per torr as, say, He.

1-2. Consider the integral

$$I^2 = \int_{-\infty}^{\infty} e^{-x^2} dx \int_{-\infty}^{\infty} e^{-y^2} dy = \iint_{-\infty}^{\infty} e^{-(x^2+y^2)} dx dy$$

in a two-dimensional space. Transforming to cylindrical coordinates, we have

$$\begin{aligned} I^2 &= \iint e^{-r^2} r dr d\phi = 2\pi \int_0^{\infty} e^{-r^2} r dr \\ &= \pi \int_0^{\infty} e^{-r^2} d(r^2) = -\pi e^{-r^2} \Big|_0^{\infty} = \pi \end{aligned}$$

Hence,

$$I = \int_{-\infty}^{\infty} e^{-x^2} dx = \sqrt{\pi}$$

Introduction To Plasma Physics Solution Manual

Julia Schneider



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Plasma Physics Richard Fitzpatrick, 2014-08-01 Encompasses the Lectured Works of a Renowned Expert in the Field Plasma Physics An Introduction is based on a series of university course lectures by a leading name in the field and thoroughly covers the physics of the fourth state of matter This book looks at non relativistic fully ionized nondegenerate quasi neutral and weakly coupled plasm

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