

### 7B-3 : Entropy Change of an Isobaric Process

6 pts

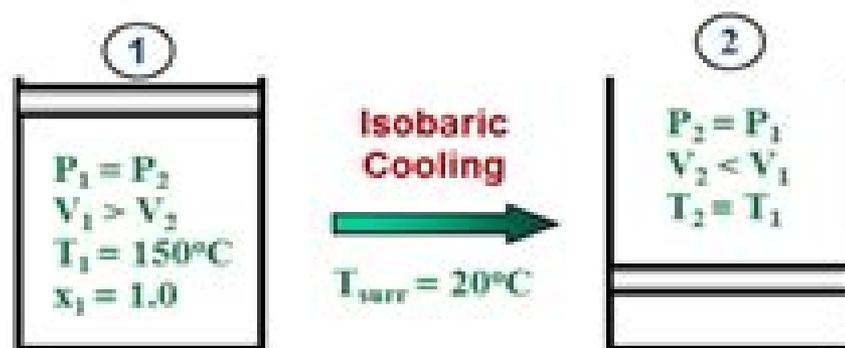
A piston-and-cylinder device with a free-floating piston contains 2.6 kg of saturated water vapor at 150°C. The water loses heat to the surroundings until the cylinder contains saturated liquid water.

The surroundings are at 20°C. Calculate... a.)  $\Delta S_{\text{water}}$ , b.)  $\Delta S_{\text{surroundings}}$ , c.)  $\Delta S_{\text{universe}}$

**Read :** Calculating  $\Delta S$  for the water in the cylinder is straightforward. The key to calculating  $\Delta S_{\text{surr}}$  is the fact that the surroundings behave as a thermal reservoir. The temperature of the surroundings does not change and there are no irreversibilities in the surroundings that are associated with the process. The key to calculating  $\Delta S_{\text{univ}}$  is the fact that the universe is made up of the combination of the system and the surroundings. Consequently,  $\Delta S_{\text{univ}} = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$ .

<b>Given:</b>	$m$	2.60	kg	<b>Find:</b>	$\Delta S_{\text{sys}}$	???	kJ/K
	$T_1$	150	°C		$\Delta S_{\text{surr}}$	???	kJ/K
	$x_1$	1	kg vap/kg		$\Delta S_{\text{univ}}$	???	kJ/K
	$T_{\text{surr}}$	20	°C				
	$P_1 = P_2$						
	$x_2$	0	kg vap/kg				

**Diagram:**



- Assumptions:**
1. Changes in kinetic and potential energies are negligible.
  2. Boundary work is the only form of work that crosses the system boundary.
  3. The surroundings behave as a thermal reservoir.

# Entropy Problems And Solutions

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## **Table of Contents Entropy Problems And Solutions**

1. Understanding the eBook Entropy Problems And Solutions
  - The Rise of Digital Reading Entropy Problems And Solutions
  - Advantages of eBooks Over Traditional Books
2. Identifying Entropy Problems And Solutions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Entropy Problems And Solutions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Entropy Problems And Solutions
  - Personalized Recommendations
  - Entropy Problems And Solutions User Reviews and Ratings
  - Entropy Problems And Solutions and Bestseller Lists
5. Accessing Entropy Problems And Solutions Free and Paid eBooks

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  - Entropy Problems And Solutions Budget-Friendly Options
6. Navigating Entropy Problems And Solutions eBook Formats
    - ePub, PDF, MOBI, and More
    - Entropy Problems And Solutions Compatibility with Devices
    - Entropy Problems And Solutions Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Entropy Problems And Solutions
    - Highlighting and Note-Taking Entropy Problems And Solutions
    - Interactive Elements Entropy Problems And Solutions
  8. Staying Engaged with Entropy Problems And Solutions
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Entropy Problems And Solutions
  9. Balancing eBooks and Physical Books Entropy Problems And Solutions
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Entropy Problems And Solutions
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Entropy Problems And Solutions
    - Setting Reading Goals Entropy Problems And Solutions
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Entropy Problems And Solutions
    - Fact-Checking eBook Content of Entropy Problems And Solutions
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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