



Student Exploration: Balancing Chemical Equations

(This activity can be completed online.)

Essential Question: How do we balance chemical equations? What is the relationship between the number of atoms of each element on the reactant side and the number of atoms of each element on the product side?

What You Will Do: You will use a simulation to balance chemical equations.

1. **Modeling:** Write a balanced chemical equation for the reaction of hydrogen gas and oxygen gas to form water. How many molecules of hydrogen gas and oxygen gas are needed to form 10 molecules of water? How many atoms of hydrogen and oxygen are in the reactants and products?
2. **Modeling:** Write a balanced chemical equation for the reaction of hydrogen gas and oxygen gas to form water. How many molecules of hydrogen gas and oxygen gas are needed to form 10 molecules of water? How many atoms of hydrogen and oxygen are in the reactants and products?

What You Will Do

Use the simulation to balance the chemical equation for the reaction of hydrogen gas and oxygen gas to form water.

Essential Question: How do we balance chemical equations? What is the relationship between the number of atoms of each element on the reactant side and the number of atoms of each element on the product side?

Procedure: Use the simulation to balance the equation.

Reaction: $2H_2 + O_2 \rightarrow 2H_2O$

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Explore Learning Student Exploration Stoichiometry Answers

Padhraic Smyth



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