

## CHAPTER 4 The Structure of Matter

## SECTION

## 2

## Ionic and Covalent Bonding

**KEY IDEAS**

As you read this section, keep these questions in mind:

- Why do atoms form bonds?
- How do ionic bonds and covalent bonds differ?
- What gives metals their distinctive properties?
- What makes a polyatomic ion different from other ions?

**Why Do Chemical Bonds Form?**

Atoms form bonds when their valence electrons interact. Recall that atoms with filled outermost energy levels are more stable than atoms with partially filled energy levels. In general, atoms join to form bonds so that each atom has a stable electron configuration. In other words, each atom binds in order to fill its outermost energy levels. ☐

There are two basic kinds of chemical bonding: ionic bonding and covalent bonding. The type of bonding in a compound determines many of the properties of the compound.

<b>A Comparison of Ionic and Covalent Compounds</b>		
	<b>Ionic compounds</b>	<b>Covalent compounds</b>
<b>Structure</b>	network of bonded ions	molecules
<b>Valence electrons</b>	transferred	shared
<b>Electrical conductivity</b>	good (when melted or dissolved)	poor
<b>State at room temperature</b>	solid	solid, liquid, or gas
<b>Melting and boiling points</b>	generally high	generally low

**What Are the Properties of Ionic Bonds?**

**Ionic bonds** form between oppositely charged ions. In general, atoms of metals, such as sodium and calcium, form positively charged ions. Atoms of nonmetals, such as chlorine and oxygen, form negatively charged ions. The attraction between ions with opposite charges holds ionic compounds together. ☐

**READING TOOLBOX**

**Compare** As you read, make a chart that describes and compares the different types of bonds.

**READING CHECK**

1. **Explain** Why do atoms join to form bonds?

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**LOOKING CLOSER**

2. **Compare** How do the structures of ionic and covalent compounds differ?

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**READING CHECK**

3. **Identify** What holds ionic compounds together?

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# Chapter Section 2 Ionic And Covalent Bonding

**Mark Steyvers**



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