

WORKSHEET
CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1
 Molecular formula of hydrocarbon or a mixture of hydrocarbons

ANSWERS KEY

QUESTION 1. Calculate the empirical formula of hydrocarbon and its molecular formula.

1. A hydrocarbon of 0.100 g is analysed and found to give 0.330 g carbon dioxide and 0.090 g water. **Find the empirical formula of hydrocarbon and its molecular formula if its relative molecular mass is 100.**
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) calculate the moles of CO_2 and H_2O and deduce the moles of C and H in 0.100 g of hydrocarbon
 $n(CO_2) = \frac{0.330}{44} = 0.0075 \text{ mol}$
 $n(H_2O) = \frac{0.090}{18} = 0.005 \text{ mol}$
 - (c) deduce the moles of C and H in 0.100 g of hydrocarbon
 $n(C) = 0.0075 \text{ mol}$
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2. A hydrocarbon of 0.100 g is analysed and found to give 0.330 g carbon dioxide and 0.090 g water. **Find the empirical formula of hydrocarbon and its molecular formula if its relative molecular mass is 100.**
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Chem 21 Lab Answers

Christian Drosten



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