

# Chemistry Chapter 4

## Balancing chemical Equations

### Chemical Reactions

- Reactions involve chemical changes in matter resulting in new substances
- Reactions involve rearrangement and exchange of atoms to produce new molecules
  - Elements are not transmuted during a reaction

### Chemical Equations



- $\text{CH}_4$  and  $\text{O}_2$  are the reactants, and  $\text{CO}_2$  and  $\text{H}_2\text{O}$  are the products
- the (g) after the formulas tells us the state of the chemical
- the number in front of each substance tells us the numbers of those molecules in the reaction
  - called the **coefficients**

### Balancing chemical equations

- Step 1: Determine the state of the reactants and products in the chemical reaction
- Step 2: Write the unbalanced chemical equation
- Step 3: Balance the number of atoms of each element in the reactant and product side, starting with the most complicated molecule

### Balancing a Hypothetical Chemical Equation



- In this hypothetical reaction, C and D are the reactants, and X and Y are the products
- The number in front of each substance tells us the numbers of those molecules of reactants and products in the reaction
- So, if you react exactly 3 moles of C with 2 moles of D, then you will produce 5 moles of X and 2 moles of Y.
- If you reacted 6 moles of C with 4 moles of D, you will produce 10 moles of X and 4 moles of Y.

### Example: Balancing chemical equations

- Glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ ) reacts with Oxygen gas to produce gaseous carbon dioxide and water vapor



Reactants

6 C

12 H

8 O

Products

1 C

2 H

3 O

# Chemistry Chapter 4 Answers

**Daniela Niemeyer**



## Chemistry Chapter 4 Answers:

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demonstrates how introducing students to metacognition and Bloom's Taxonomy reveals to them the importance of understanding how they learn and provides the lens through which they can view learning activities and measure their intellectual growth. Next she presents a specific study system that can quickly empower students to maximize their learning. Then she addresses the importance of dealing with emotion, attitudes, and motivation by suggesting ways to change students' mindsets about ability and by providing a range of strategies to boost motivation and learning. Finally, she offers guidance to faculty on partnering with campus learning centers. She pays particular attention to academically unprepared students, noting that the strategies she offers for this particular population are equally beneficial for all students. While stressing that there are many ways to teach effectively and that readers can be flexible in picking and choosing among the strategies, she presents Sandra McGuire's step-by-step process for delivering the key messages of the book to students in as little as 50 minutes. Free online supplements provide three slide sets and a sample video lecture. This book is written primarily for faculty but will be equally useful for TAs, tutors, and learning center professionals. For readers with no background in education or cognitive psychology, the book avoids jargon and esoteric theory.

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