

Intro to Chemistry Final Exam Study Guide

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. The Fahrenheit temperature scale is defined by which of the following temperatures?
a. Water boils at 100 degrees and freezes at 32 degrees.
b. Water boils at 212 degrees and freezes at 32 degrees.
c. Water evaporates at 212 degrees and freezes at 0 degrees.
d. Liquid water turns to a gas at 100 degrees and to a solid at 0 degrees.
- _____ 2. A cold-blooded reptile basks on a warm rock in the sun. Its body is warmed by
a. radiation. c. convection.
b. conduction. d. Both (a) and (b).
- _____ 3. Temperature is a measure of the average _____ energy of an object's particles.
a. mechanical c. potential
b. kinetic d. light
- _____ 4. The temperature at which all molecular motion stops is
a. 0°C. c. 0 K.
b. 0°F. d. All of the above.
- _____ 5. Convert 500°F to degrees Celsius.
a. 260°C c. 842°C
b. 296°C d. 958°C
- _____ 6. Convert 300 K to the Celsius scale.
a. 0°C c. 73°C
b. 27°C d. 573°C
- _____ 7. Convert 458 K to the equivalent Fahrenheit temperature.
a. 185°F c. 365°F
b. 237°F d. 731°F
- _____ 8. Which of the following temperatures is impossible to measure?
a. -85°F c. -20 K
b. -50°C d. 545°F
- _____ 9. What is -175°C on the Kelvin scale?
a. 76 K c. 98 K
b. 89 K d. 448 K
- _____ 10. The principles of atomic theory recognized today were conceived by
a. Avogadro. c. Dalton.
b. Bohr. d. Rutherford.
- _____ 11. The total number of protons and neutrons in the nucleus of an atom is its
a. atomic number. c. mass number.
b. Avogadro number. d. average atomic mass.
- _____ 12. All atoms of the same element have the same
a. atomic mass. c. mass number.
b. number of neutrons. d. atomic number.
- _____ 13. Atoms of the same element can differ in
a. chemical properties. c. atomic number.
b. mass number. d. number of protons and electrons.
- _____ 14. Chlorine has atomic number 17 and mass number 35. It has
a. 17 protons, 17 electrons, and 18 neutrons.
b. 35 protons, 35 electrons, and 17 neutrons.
c. 17 protons, 17 electrons, and 52 neutrons.
d. 18 protons, 18 electrons, and 17 neutrons.
- _____ 15. Phosphorus-33 (atomic number 15) contains
a. 33 protons. c. 33 neutrons.
b. 18 neutrons. d. 18 protons.
- _____ 16. Neon-22 contains 12 neutrons. It also contains
a. 12 protons. c. 22 electrons.
b. 22 protons. d. 10 protons.
- _____ 17. Valence electrons determine an atom's

Chemistry Final Exam Study Guide Multiple Choice

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



Chemistry Final Exam Study Guide Multiple Choice:

This book delves into Chemistry Final Exam Study Guide Multiple Choice. Chemistry Final Exam Study Guide Multiple Choice is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chemistry Final Exam Study Guide Multiple Choice, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Chemistry Final Exam Study Guide Multiple Choice
- Chapter 2: Essential Elements of Chemistry Final Exam Study Guide Multiple Choice
- Chapter 3: Chemistry Final Exam Study Guide Multiple Choice in Everyday Life
- Chapter 4: Chemistry Final Exam Study Guide Multiple Choice in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Chemistry Final Exam Study Guide Multiple Choice. The first chapter will explore what Chemistry Final Exam Study Guide Multiple Choice is, why Chemistry Final Exam Study Guide Multiple Choice is vital, and how to effectively learn about Chemistry Final Exam Study Guide Multiple Choice.

3. In chapter 2, the author will delve into the foundational concepts of Chemistry Final Exam Study Guide Multiple Choice. This chapter will elucidate the essential principles that must be understood to grasp Chemistry Final Exam Study Guide Multiple Choice in its entirety.

4. In chapter 3, this book will examine the practical applications of Chemistry Final Exam Study Guide Multiple Choice in daily life. This chapter will showcase real-world examples of how Chemistry Final Exam Study Guide Multiple Choice can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Chemistry Final Exam Study Guide Multiple Choice in specific contexts. The fourth chapter will explore how Chemistry Final Exam Study Guide Multiple Choice is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Chemistry Final Exam Study Guide Multiple Choice. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Chemistry Final Exam Study Guide Multiple Choice.

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