

## Flame Test Lab

### Activity Key

#### Background

The heat from a laboratory burner will cause the ions of some elements to give off light. Electrons will absorb the heat energy from the flame and will “jump” to a higher energy level. When the electrons return to their original energy levels, this absorbed energy is released as light. Different elements absorb and release different amounts of energy and thus produce different kinds of light. This light may be visible or invisible, infrared or ultraviolet radiation. Analysis of this light can be used to identify elements and molecules. In today’s lab, we will observe the visible range of light emission.

#### Objectives

1. Perform a flame test to identify the characteristic color of metal ions.
2. Identify the metal ion in an unknown solution.
3. Calculate the energy of emitted photons.

#### Materials: (per lab group)

10 ml dilute solutions of the following:

Barium Chloride  
Calcium Chloride  
Lithium Chloride  
Potassium Chloride  
Sodium Chloride  
Strontium Chloride  
Unknown Solution

7 Wood Splints  
1 Bunsen Burner  
Safety Goggles & Aprons

#### Note:

If chloride compounds are not available, metal nitrate compounds may be substituted.

Use dilute or approximately 0.1 M solutions

#### Unknowns:

Number the beakers 1-12.  
Add solutions as follows:

- |                       |                        |
|-----------------------|------------------------|
| 1. Strontium Chloride | 7. Calcium Chloride    |
| 2. Calcium Chloride   | 8. Strontium Chloride  |
| 3. Potassium Chloride | 9. Barium Chloride     |
| 4. Barium Chloride    | 10. Potassium Chloride |
| 5. Lithium Chloride   | 11. Sodium Chloride    |
| 6. Sodium Chloride    | 12. Lithium Chloride   |

#### Procedure

1. Put on your safety goggles and an apron.
2. Obtain a wood splint and the barium chloride solution.
3. Dip the wood splint into the solution and allow it to soak for 1-2 minutes.
4. Light Bunsen Burner. USE CAUTION!
5. Remove the splint from the solution and hold the end of the splint into the flame. Do not burn the splint.
6. Observe and record the color of the flame.
7. Repeat steps 2-6 using the remaining chloride solutions.
8. Obtain an unknown solution from your teacher. Record the number of your unknown.
9. Soak a wood splint in the unknown solution.
10. Hold the splint in the flame and record the color of the flame that is produced.
11. Using your data, identify the metal ion in your unknown solution.

# Chemistry Flame Test Lab Answers

**Emilie Sanchez**



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## **Table of Contents Chemistry Flame Test Lab Answers**

1. Understanding the eBook Chemistry Flame Test Lab Answers
  - The Rise of Digital Reading Chemistry Flame Test Lab Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Flame Test Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Flame Test Lab Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Flame Test Lab Answers
  - Personalized Recommendations
  - Chemistry Flame Test Lab Answers User Reviews and Ratings
  - Chemistry Flame Test Lab Answers and Bestseller Lists
5. Accessing Chemistry Flame Test Lab Answers Free and Paid eBooks
  - Chemistry Flame Test Lab Answers Public Domain eBooks
  - Chemistry Flame Test Lab Answers eBook Subscription Services

- Chemistry Flame Test Lab Answers Budget-Friendly Options
- 6. Navigating Chemistry Flame Test Lab Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Flame Test Lab Answers Compatibility with Devices
  - Chemistry Flame Test Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Flame Test Lab Answers
  - Highlighting and Note-Taking Chemistry Flame Test Lab Answers
  - Interactive Elements Chemistry Flame Test Lab Answers
- 8. Staying Engaged with Chemistry Flame Test Lab Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Flame Test Lab Answers
- 9. Balancing eBooks and Physical Books Chemistry Flame Test Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Flame Test Lab Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Flame Test Lab Answers
  - Setting Reading Goals Chemistry Flame Test Lab Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Flame Test Lab Answers
  - Fact-Checking eBook Content of Chemistry Flame Test Lab Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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