

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

### Periodic Trends Practice Worksheet

1. What is the difference between electronegativity and ionization energy?
2. Why does fluorine have a higher ionization energy than iodine?
3. Why do elements in the same family generally have similar properties?
4. Which is the largest atom in Group 4a (14)?
5. Which is the smallest atom in Group 7a (17)?
6. Which is the smallest atom in period 5?
7. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
8. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
9. Arrange the following atoms in order of decreasing atomic radius. Na Al P Cl Mg
10. For each of the following pairs, circle the element that is larger?  
a)  $N^{3-}$  or  $F^{-1}$       b)  $Mg^{2+}$  or  $Ca^{2+}$       c)  $Fe^{2+}$  or  $Fe^{3+}$
11. Circle the element in each pair has the larger radius?  
a) Mg or  $Mg^{2+}$       b) O or  $O^{2-}$       c)  $K^{+}$  or Cl      d)  $P^{3-}$  or  $S^{2-}$
12. In each of the following pairs, circle the species with the higher first ionization energy:  
a) Li or Cs      (b) Cl or Ar      (c) Ca or Br      (d) Na or Ne      (e) B or Be
13. In each of the following pairs, circle the species with the larger atomic radius:  
a) Mg or Ba      (b) S or  $S^{2-}$       (c)  $Cu^{+2}$  or Cu      (d) He or H      (e) Na or Cl
14. Circle the best choice in each list:  
a) highest first ionization energy: C, N, Si  
b) highest electronegativity: As, Sn, S  
c) largest radius:  $S^{2-}$ ,  $Cl^{-}$ , Cl
15. Order the following groups from largest to smallest radii.  
a) Ar, Cl, K,  $S^{2-}$   
b) C, Al, F, Si  
c) Na, Mg, Ar, P  
d)  $I^{-}$ ,  $Ba^{2+}$ ,  $Cs^{+1}$ ,  $F^{-1}$

# Chemistry Practice Test Periodic Trends And Orbitals

**S Marginson**



## **Chemistry Practice Test Periodic Trends And Orbitals:**

## Unveiling the Magic of Words: A Report on "**Chemistry Practice Test Periodic Trends And Orbitals**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Chemistry Practice Test Periodic Trends And Orbitals**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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