

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative	noble gas	octet	electron configuration
positive	electrons	ions	polyatomic
crisscross	anion	cation	empirical formula

- An ionic bond is an attraction between oppositely charged ions.
- Anions have a negative charge.
- An atom becomes an ion by losing or gaining electrons.
- The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
- When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
- The atoms that make up polyatomic ions are bonded together by covalent bonds.
- The crisscross method can be used to write the formula for an ionic compound.
- The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- | | |
|------------------------|--|
| <u>losing</u> | 9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons. |
| <u>true</u> | 10. A cation has a <u>positive</u> charge. |
| <u>true</u> | 11. A binary ionic compound contains only <u>one</u> kind of cation and one kind of anion. |
| <u>polyatomic</u> | 12. <u>Monatomic</u> ions consist of more than one atom. |
| <u>seven</u> | 13. The Lewis dot diagram for chlorine, a group 7A element, has <u>six</u> electrons. |
| <u>true (released)</u> | 14. A <u>great deal of energy is produced</u> when an ionic compound is formed. |

Chemistry Review And Reinforcement Answers

M Mark



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Table of Contents Chemistry Review And Reinforcement Answers

1. Understanding the eBook Chemistry Review And Reinforcement Answers
 - The Rise of Digital Reading Chemistry Review And Reinforcement Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Review And Reinforcement Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Review And Reinforcement Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Review And Reinforcement Answers
 - Personalized Recommendations
 - Chemistry Review And Reinforcement Answers User Reviews and Ratings
 - Chemistry Review And Reinforcement Answers and Bestseller Lists
5. Accessing Chemistry Review And Reinforcement Answers Free and Paid eBooks
 - Chemistry Review And Reinforcement Answers Public Domain eBooks
 - Chemistry Review And Reinforcement Answers eBook Subscription Services

- Chemistry Review And Reinforcement Answers Budget-Friendly Options
- 6. Navigating Chemistry Review And Reinforcement Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Review And Reinforcement Answers Compatibility with Devices
 - Chemistry Review And Reinforcement Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Review And Reinforcement Answers
 - Highlighting and Note-Taking Chemistry Review And Reinforcement Answers
 - Interactive Elements Chemistry Review And Reinforcement Answers
- 8. Staying Engaged with Chemistry Review And Reinforcement Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Review And Reinforcement Answers
- 9. Balancing eBooks and Physical Books Chemistry Review And Reinforcement Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Review And Reinforcement Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Review And Reinforcement Answers
 - Setting Reading Goals Chemistry Review And Reinforcement Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Review And Reinforcement Answers
 - Fact-Checking eBook Content of Chemistry Review And Reinforcement Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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