

# Solutions

## Exercises

1. Define the term solution. How many types of solutions are formed? Write briefly about each type with an example.

**Ans.** A solution is a homogeneous mixture of two or more pure substances. Solute and solvent are two components of a solution.

**Types of solutions:** On the basis of different states of solute and solvent solutions are of three types mainly, which are further categorized into three.

Type of Solution	Solute	Solvent	Common Examples
Gaseous Solutions	Gas	Gas	Mixture of oxygen and nitrogen gases
	Liquid	Gas	Chloroform mixed with nitrogen gas
	Solid	Gas	Camphor in nitrogen gas
Liquid Solutions	Gas	Liquid	Oxygen dissolved in water
	Liquid	Liquid	Ethanol dissolved in water
	Solid	Liquid	Glucose dissolved in water
Solid Solutions	Gas	Solid	Solution of hydrogen in palladium
	Liquid	Solid	Amalgam of mercury with sodium
	Solid	Solid	Copper dissolved in gold

2. Give an example of a solid solution in which the solute is a gas.

**Ans.** In case a solid solution is formed between two substances, an interstitial solid solution will be formed. For example: a solution of hydrogen in palladium is a solution in which the solute is a gas.

3. Define the following terms:

(i) Mole fraction

(iii) Molarity

(ii) Molality

(iv) Mass percentage

**Ans.** (i) **Mole fraction:** It is defined as the ratio of the number of moles of the individual component (solute or solvent) to the total number of moles of all components present in the solution. If  $n_A$  is the number of moles of solute dissolved in  $n_B$  moles of solvent, then Mole fraction of solute

$$(X_A) = \frac{n_A}{n_A + n_B}$$

$$\text{Mole fraction of solvent } (X_B) = \frac{n_B}{n_A + n_B}$$

(ii) **Molality:** It is defined as the number of moles of a solute present in 1000g (1kg) of a solvent.

$$\text{Molality } (m) = \frac{\text{Number of moles of solute}}{\text{Weight of solvent in kg}} = \frac{n}{W}$$

$$\text{Unit} = \text{mol kg}^{-1} \text{ or m(molal)}$$

(iii) **Molarity:** It is defined as the number of moles of solute present in one litre of solution.

$$\text{Molarity } (M) = \frac{\text{Number of moles of solute}}{\text{Volume of solution in litre}} = \frac{n}{V}$$

$$\text{Unit} = \text{mol L}^{-1} \text{ or M(Molar)}$$

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**Kevin Smith**



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**Chemistry (Teacher Guide)** Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books. The teacher is one who guides students through the subject matter, helps each student stay on schedule and be organized, and is their source of accountability along the way. With that in mind, this guide provides additional help through the laboratory exercises as well as lessons, quizzes, and examinations that are provided along with the answers. The lessons in this study emphasize working through procedures and problem solving by learning patterns. The vocabulary is kept at the essential level. Practice exercises are given with their answers so that the patterns can be used in problem solving. These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college. Guided labs are provided to enhance instruction of weekly lessons. There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions. It is important to see the hand of God and His principles and wisdom as it plays out in chemistry. This course integrates what God has told us in the context of this study. Features: Each suggested weekly schedule has five easy-to-manage lessons that combine reading and worksheets. Worksheets, quizzes, and tests are perforated and three-hole punched; materials are easy to tear out, hand out, grade, and store. Adjust the schedule and materials needed to best work within your educational program. Space is given for assignments, dates. There is flexibility in scheduling. Adapt the days to your school schedule. Workflow: Students will read the pages in their book and then complete each section of the teacher guide. They should be encouraged to complete as many of the activities and projects as possible as well. Tests are given at regular intervals with space to record each grade. About the Author: DR DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and Au Sable Institute of Environmental Studies. *Publication, 1965 Student Solutions Manual for Gustafson and Frisk's Beginning Algebra* Diane Koenig, 1992

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**Teacher's Manual for First-year Mathematics** George William Myers,1911

**AutoCAD and Its Applications** Terence M. Shumaker,David A. Madsen,2000 These editions of AutoCAD and its Applications provide instruction for mastering AutoCAD RM 2000 commands and drawing and dimensioning techniques Like the Release 14 editions the AutoCAD 2000 title offers a Basics and Advanced edition This allows for manageable texts in both size and content as well as flexibility to meet the needs of various course structures Content of the Basics edition provides comprehensive coverage of introductory and two dimensional AutoCAD drafting while the Advanced edition covers three dimensional and other advanced functions Both texts cover topics in an easy to understand sequence and progress in a manner that allows students to become comfortable with AutoCAD In depth discussions of every major new and existing AutoCAD feature command and option are provided Hundreds of exercises questions and drawing problems assist learning No AutoCAD book surpasses the depth of coverage provided by this outstanding title Command initiation methods appear in the text margin next to command introduction Professional Tips explain how to use AutoCAD efficiently Fold out cover illustrates AutoCAD command buttons and screen and pull down menu trees Topics keyed to AutoCAD User Guide Drawing Problems and Exercises offer application to several disciplines clearly identified by an icon Chapter Tests allow review of important commands and concepts

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sketches and plots Contains large number of examples and exercises Provides solutions to all exercises Includes R codes  
Silvelyn Zwanzig is a Professor for Mathematical Statistics at Uppsala University She studied Mathematics at the Humboldt University of Berlin Before coming to Sweden she was Assistant Professor at the University of Hamburg in Germany She received her Ph D in Mathematics at the Academy of Sciences of the GDR She has taught Statistics to undergraduate and graduate students since 1991 Her research interests include theoretical statistics and computer intensive methods Rauf Ahmad is Associate Professor at the Department of Statistics Uppsala University He did his Ph D at the University of Gttingen Germany Before joining Uppsala University he worked at the Division of Mathematical Statistics Department of Mathematics Link ping University and at Biometry Division Swedish University of Agricultural Sciences Uppsala He has taught Statistics to undergraduate and graduate students since 1995 His research interests include high dimensional inference mathematical statistics and U statistics Special Assessment Procedure Arthur Rowland Burnstan,1929

**Special Report** New York (State). State Tax Commission,1929 *A Manual of Arithmetic for Teachers, Giving Practical Suggestions for Teaching of Arithmetic in the Elementary Grades and Junior High School* Edison Ellsworth Oberholtzer,1920

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