

Section 1.1

Section 1.1

Section 1.1 covers the following topics in the areas provided:

1. The structure of a chemical compound depends on:
 - (a) the number of atoms of each element and protons
 - (b) the relative number of neutrons of isotopes and protons
 - (c) number of atoms of each element in each compound
 - (d) number of valence electrons, electronegativity, etc.
2. The structure of a chemical compound depends on:
 - (a) the number of atoms of each element and protons
 - (b) the number of valence electrons of each element and protons
 - (c) the number of neutrons of isotopes that combine with a certain quantity of protons
 - (d) the relative electronegativity of each element (e.g., C, O, N, H, etc.)
3. The structure of a chemical compound depends on:
 - (a) the number of atoms of each element and protons
 - (b) the energy released in the reaction
 - (c) relative electronegativity of all elements in the reaction
 - (d) number of valence electrons of each element
4. The structure of a chemical compound depends on:
 - (a) the number of atoms of each element and protons
 - (b) the number of valence electrons
 - (c) electronegativity
 - (d) relative electronegativity

Section 1.1 covers the following topics in the areas provided:

$$2H_2 + O_2 \rightarrow 2H_2O$$

a. Balance the equation:



b. How many grams of H_2 are needed to react with 16g of O_2 ?

1. The mass of H_2 is _____ g. The mass of O_2 is _____ g. The mass of H_2O is _____ g.

2. The mass of H_2 is _____ g. The mass of O_2 is _____ g. The mass of H_2O is _____ g.

3. The mass of H_2 is _____ g. The mass of O_2 is _____ g. The mass of H_2O is _____ g.

Chapter 9 Review Stoichiometry Section 3

Anthony S. Fauci



Chapter 9 Review Stoichiometry Section 3:

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