

Exam-style questions and sample answers have been written by the author. In examinations, the way marks are awarded may be different.

# Teacher's resource answers

## Chapter 15

1 a Isomers are compounds with the same molecular formula [1]

but different structural formulae. [1]

b [2]



2 a heat change = mass  $\times$  specific heat capacity  $\times$  change in temperature [1]  
 $= 200 \times 4.18 \times 30$  [1]  
 $= 25\,080 \text{ J} = 25.1 \text{ kJ}$  [1]

b number of moles of 2-methylbutane [1]  
 $= \frac{0.626}{72} = 8.69 \times 10^{-3} \text{ mol}$  [1]

standard enthalpy of combustion [1]  
 $= \frac{25.1}{8.69 \times 10^{-3}}$  [1]

$= -2890 \text{ kJ mol}^{-1}$  [1]

c heat change is the same, 25.1 kJ [1]  
 number of moles of 2,2-dimethylpropane

$= \frac{171}{24000} = 7.13 \times 10^{-3}$  [1]

standard enthalpy of combustion [1]  
 $= \frac{25.1}{7.13 \times 10^{-3}}$  [1]

$= -3520 \text{ kJ mol}^{-1}$  [1]



cycle with carbon dioxide and water as the third corner and where the directions of the three arrows are correct, and [1]

labelling arrows with correct enthalpies [1]

$\Delta H_{\text{combustion}} + (-3520) = -2890$  [1]

$\Delta H_{\text{combustion}} = -2890 - (-3520) = +630 \text{ kJ mol}^{-1}$  [1]

4 The 2-methylbutane has a greater surface area. [1]

therefore a greater number of van der Waals' forces/induced dipoles. [1]

It has stronger intermolecular forces and therefore more energy required to separate molecules (and a higher boiling point). [1]



1 mark for each structure

Note: Carbon and hydrogen atoms are shown for displayed formulae. [2]

6 a  $\text{C}_2\text{H}_6 + \text{Br}_2 \rightarrow \text{C}_2\text{H}_5\text{Br} + \text{HBr}$  [1]  
 correct reactants [1]

for correct products [1]

b Initiation. [1]

# Core Teaching Resources Chemistry Answers Chapter 15

**M Mark**



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## **Table of Contents Core Teaching Resources Chemistry Answers Chapter 15**

1. Understanding the eBook Core Teaching Resources Chemistry Answers Chapter 15
  - The Rise of Digital Reading Core Teaching Resources Chemistry Answers Chapter 15
  - Advantages of eBooks Over Traditional Books
2. Identifying Core Teaching Resources Chemistry Answers Chapter 15
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Core Teaching Resources Chemistry Answers Chapter 15
  - User-Friendly Interface
4. Exploring eBook Recommendations from Core Teaching Resources Chemistry Answers Chapter 15
  - Personalized Recommendations
  - Core Teaching Resources Chemistry Answers Chapter 15 User Reviews and Ratings
  - Core Teaching Resources Chemistry Answers Chapter 15 and Bestseller Lists

5. Accessing Core Teaching Resources Chemistry Answers Chapter 15 Free and Paid eBooks
  - Core Teaching Resources Chemistry Answers Chapter 15 Public Domain eBooks
  - Core Teaching Resources Chemistry Answers Chapter 15 eBook Subscription Services
  - Core Teaching Resources Chemistry Answers Chapter 15 Budget-Friendly Options
6. Navigating Core Teaching Resources Chemistry Answers Chapter 15 eBook Formats
  - ePub, PDF, MOBI, and More
  - Core Teaching Resources Chemistry Answers Chapter 15 Compatibility with Devices
  - Core Teaching Resources Chemistry Answers Chapter 15 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Core Teaching Resources Chemistry Answers Chapter 15
  - Highlighting and Note-Taking Core Teaching Resources Chemistry Answers Chapter 15
  - Interactive Elements Core Teaching Resources Chemistry Answers Chapter 15
8. Staying Engaged with Core Teaching Resources Chemistry Answers Chapter 15
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Core Teaching Resources Chemistry Answers Chapter 15
9. Balancing eBooks and Physical Books Core Teaching Resources Chemistry Answers Chapter 15
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Core Teaching Resources Chemistry Answers Chapter 15
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Core Teaching Resources Chemistry Answers Chapter 15
  - Setting Reading Goals Core Teaching Resources Chemistry Answers Chapter 15
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Core Teaching Resources Chemistry Answers Chapter 15
  - Fact-Checking eBook Content of Core Teaching Resources Chemistry Answers Chapter 15
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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