



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or covalent.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{O}_2$
- Write the formula for the following compounds:
  - hydrogen chloride
  - hydrogen sulfide
  - hydrogen cyanide
  - hydrogen peroxide
  - hydrogen selenide
  - hydrogen telluride

- Write the formula for the following compounds:
  - hydrogen bromide
  - hydrogen iodide
  - hydrogen fluoride
  - hydrogen phosphide
  - hydrogen selenide
  - hydrogen telluride
- Write the formula for the following compounds:
  - hydrogen chloride
  - hydrogen sulfide
  - hydrogen cyanide
  - hydrogen peroxide
  - hydrogen selenide
  - hydrogen telluride

### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride.  $\text{H}-\text{F}$
- Draw the Lewis structure for hydrogen chloride.  $\text{H}-\text{Cl}$
- Draw the Lewis structure for hydrogen bromide.  $\text{H}-\text{Br}$
- Draw the Lewis structure for hydrogen iodide.  $\text{H}-\text{I}$
- Draw the Lewis structure for hydrogen cyanide.  $\text{H}-\text{C}\equiv\text{N}$
- Draw the Lewis structure for hydrogen sulfide.  $\text{H}-\text{S}-\text{H}$
- Draw the Lewis structure for hydrogen selenide.  $\text{H}-\text{Se}-\text{H}$
- Draw the Lewis structure for hydrogen telluride.  $\text{H}-\text{Te}-\text{H}$



### SECTION 8.3 BONDING IN THE CARBON

- Write the Lewis structure for ethane.  $\text{C}_2\text{H}_6$
- Write the Lewis structure for ethene.  $\text{C}_2\text{H}_4$
- Write the Lewis structure for ethyne.  $\text{C}_2\text{H}_2$

# Covalent Bonding Pearson Answers

**JG Myers**



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