

### 3. Atoms, elements and compounds



Atomic mass = 23

Proton number = 11

Neutron number = 12 (23 - 11)

mass of electron =  $\frac{1}{1836}$

**ISOTOPE**: atoms of the same element with the same proton number but a different neutron number

- Have the same properties due to the same number of electrons on the outer shell

**Radioisotope**

- testing for leaks in pipes
- treating cancer

#### IONIC BONDING

**Bonding**: oppositely charged ions are strongly attracted by electrostatic force + involves electron transfer

metal + non metal

- metals give electrons to non-metals

**Properties**: strong bonds due to attractive forces

- conduct electricity
- high mp/bp
- charged particles
- soluble in water

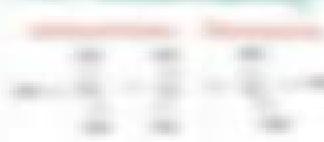
#### Macromolecular compounds

**SILICON (IV) OXIDE**

**Properties**



#### Covalent bonding



... the strong nature of covalent bonding means that ...

A large box containing a list of properties and examples of covalent compounds. Properties include low melting and boiling points, poor conductors of electricity, and solubility in water. Examples include water, carbon dioxide, and methane.

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