

Question 1

1(a) Balance (top loading), spatula, stirring rod, dropper, pipette (10 cm³), pipette (25 cm³), glass (filter) funnel, burette.

1(b)

Burette readings (cm ³)	Titration 1	Titration 2	Titration 3
Final volume	2.00	1.85	1.50
Initial volume	21.95	21.85	19.50
Volume of solution used	19.95*	20.00*	18.00

1(c)(i) See table above.

1(c)(ii) Average volume = (volume used in titration 1 + volume used in titration 2)/2.

Average volume = $\{(19.95 + 20.00)/2\}$ cm³. i.e. 19.98 cm³.

Actual accurate value to be used in subsequent calculations is 19.975 cm³.

1(d) 1000 cm³ K₂Cr₂O_{7(aq)} contains 0.020 mol. K₂Cr₂O₇. 19.98 cm³ K₂Cr₂O_{7(aq)} contains $(19.98 \times (0.020/1000))$ mol. K₂Cr₂O₇, i.e. 4.00×10^{-4} mol K₂Cr₂O₇.

Actual accurate value to be used in subsequent calculations is 3.995×10^{-4} mol K₂Cr₂O₇.

1(e) From the balanced equation, the Fe²⁺:Cr₂O₇²⁻ stoichiometric ratio is 6:1. So 6 moles of iron ions (Fe²⁺) react with one mole of dichromate ions (Cr₂O₇²⁻).

1(f) Moles of iron in sample = (6 × moles of Cr₂O₇²⁻ used in titration) i.e. $(6 \times 4.00 \times 10^{-4})$ mol. = 2.4×10^{-3} mol. iron.

Actual accurate value to be used in subsequent calculations is 2.397×10^{-3} mol.

1(g) Mass of iron in salt sample = (moles of iron in sample × molar mass of iron).

Mass of iron in salt sample = $(2.4 \times 10^{-3} \times 55.8)$ g iron i.e. 0.134 g iron.

Actual accurate value to be used in subsequent calculations is 0.1337526 g.

1(h) Percentage of iron in iron salt sample = (mass of iron in sample/mass of sample) × 100%.

Percentage of iron in iron salt sample = $(0.134/0.500) \times 100\%$ i.e. 26.80% iron.

If accurate values were used in all calculation steps leading to this one, the final answer would be 26.75%.

1(i) Reduction of a species is the gain of an electron, or electrons, by that species.

1(j) A reducing agent is a species which causes the reduction of another species, and is itself oxidized when it reacts; it loses an electron, or electrons.

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