
UNE Medical Biochemistry Exam 1 Questions With Complete Solutions.

what percentage of water is in the body?

60%

Relationship between protons and acids & bases

Acids donate proton. Bases accept proton

Strong Acids Vs Weak Acids

Strong acids fully dissociate (100%) and weak acids partially dissociate

properties of water

dipole dipole

polar

H bonding

dissolves electrolytes (CL /Na)

Small K_d - doesn't dissociate alot

kw ion product

10^{-14}

pH

$-\log[H^+]$

biological PH

7.4

types of buffers

phosphates, bicarbonate, proteins

Buffers

weak acid and conjugate base

K_a

how well acid can dissociate

pka

$-\log(K_a)$

50% dissociation

Buffers work best when?

$pH = pKa$

within +/- 1 pH unit of pka

Henderson-Hasselbalch equation

$pH = pKa + \log \frac{A^-}{HA}$

$pH < pKa$

protonated

$pH > pKa$

ionized

amino acids are linked together by

peptide bonds between carboxyl and amino group

what's special about histidine pka?

close to biological PH of 7.4 so it can act as an donor or acceptor

nonpolar amino acids

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TD Snyder



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