

LSA =  $2\pi r h$   
 TSA =  $2\pi r^2 + 2\pi r h$

5. Cone:

$V = \frac{1}{3} \pi r^2 h$

LSA =  $\pi r \sqrt{r^2 + h^2}$

TSA =  $\pi r^2 + \pi r \sqrt{r^2 + h^2}$



6. Spherical Segment:

$V = \frac{\pi h^2}{3} (3R - h)$

Z = area of zone =  $2\pi R h$



7. Pyramid:

$V = \frac{1}{3} A h$



8. Frustums:

a. Cone:

$V = \frac{\pi h}{3} (R^2 + Rr + r^2)$



b. Pyramid:

$V = \frac{1}{3} h (A_1 + \sqrt{A_1 A_2} + A_2)$



# Mensuration Problems With Solutions

**K Morrison**

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web the pcl<sub>3</sub> lewis structure has the typical case of phosphorus p in the center with 3 bonds to 3 other atoms phosphorus is from the same column as nitrogen in the periodic table meaning that p and n generally have the same bonding structure note the lone pair dots without bonds on top of p just like for n in the previous example for nh<sub>3</sub> chlorine cl

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