

## Chemistry – Unit 4 Notes Dalton's Playbook

In the late 19<sup>th</sup> century, British physicist John Dalton developed and refined an atomic model which explained experimental observations. Dalton's theory of atoms and molecules included several ideas. The substance on which Dalton's theory is based is the atom. It is a smallest particle which can't be created or destroyed. Atoms are indivisible particles which cannot be created or destroyed of their own.

### Table 1 – Properties

Mass	Volume	Shape	Intermolecular force
Mass is conserved			
Volume is not			

1. What happens to the mass of the matter as it goes from solid?
2. What happens to the volume of the matter as it goes from solid to liquid?
3. What are the intermolecular forces in solid, liquid and gas states of the matter as it goes from solid.

### Table 2 – Examples

This table need to record the actual and theoretical percentage composition of each of the compounds mentioned in the table.

Compound	Actual %	Theoretical %	Example
Mass composition			
Volume composition			
Mass of element			
Atomic weight of element			

1. Write equations for the formation of the gases in laboratory conditions and compare your data.
2. How do the masses of the gas in all the experiments follow Dalton's theory? Do you have any other examples?

# Modeling Chemistry Dalton Playhouse Notes Answers

**Patrick Vollmar**



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web these requirements are based on decades of industry knowledge and research by the coffee brewing center all sca certified brewers have met these requirements which

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web for example the scaa brewing standard for preparation of drip coffee defines the proper ratios of water to coffee the proper extraction brewing temperature and holding

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web the ratio of coffee to water when preparing a brew in this standard brew ratio is expressed in grams of coffee per 1 000 kg of fresh water at 17 3 c g kg the brew

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