

# Atomic structure and the periodic table

## Atomic structure



protons  
neutrons  
electrons

atoms have no overall charge

particle	relative mass	relative charge
proton	1	+1
neutron	1	0
electron	very small	-1

## elements

elements - substances made up of atoms with the same atomic number

isotopes of an element - atoms with the same number of protons but different numbers of neutrons

relative atomic mass - the average mass number for an element

## Equations, mixtures and chromatography

### chemical formulas and equations

chemical formula - shows the proportion of atoms of each element in a compound

chemical equation - shows the overall change in a reaction

example:

reactants

products

word equation: methane + oxygen → carbon dioxide + water

symbol equation:  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

chromatogram

insoluble components stay on the base line

components separate as they move up the paper

## nuclear symbols

mass number = total number of protons and neutrons in an atom



atomic number = number of protons in an atom

## compounds

compound - a substance formed from two or more elements chemically bonded together in fixed proportions

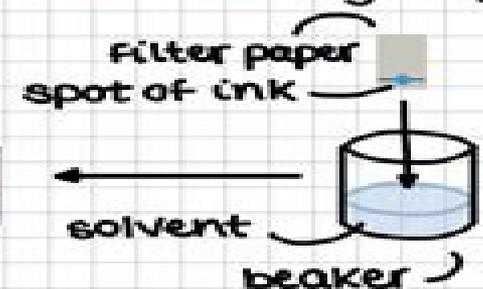
molecule - particle containing two or more non-metal atoms bonded covalently. Molecules can be elements or compounds

## mixtures

mixtures - substances made up of different elements or compounds that are not chemically bonded to each other

mixtures can be separated using physical methods

## paper chromatography



# Atomic Structure Guided Notes Answers

**Lingjun Ying**



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## **Table of Contents Atomic Structure Guided Notes Answers**

1. Understanding the eBook Atomic Structure Guided Notes Answers
  - The Rise of Digital Reading Atomic Structure Guided Notes Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Guided Notes Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Atomic Structure Guided Notes Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Guided Notes Answers
  - Personalized Recommendations
  - Atomic Structure Guided Notes Answers User Reviews and Ratings
  - Atomic Structure Guided Notes Answers and Bestseller Lists
5. Accessing Atomic Structure Guided Notes Answers Free and Paid eBooks
  - Atomic Structure Guided Notes Answers Public Domain eBooks
  - Atomic Structure Guided Notes Answers eBook Subscription Services

- Atomic Structure Guided Notes Answers Budget-Friendly Options
- 6. Navigating Atomic Structure Guided Notes Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Atomic Structure Guided Notes Answers Compatibility with Devices
  - Atomic Structure Guided Notes Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Atomic Structure Guided Notes Answers
  - Highlighting and Note-Taking Atomic Structure Guided Notes Answers
  - Interactive Elements Atomic Structure Guided Notes Answers
- 8. Staying Engaged with Atomic Structure Guided Notes Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Atomic Structure Guided Notes Answers
- 9. Balancing eBooks and Physical Books Atomic Structure Guided Notes Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Atomic Structure Guided Notes Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Atomic Structure Guided Notes Answers
  - Setting Reading Goals Atomic Structure Guided Notes Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Atomic Structure Guided Notes Answers
  - Fact-Checking eBook Content of Atomic Structure Guided Notes Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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