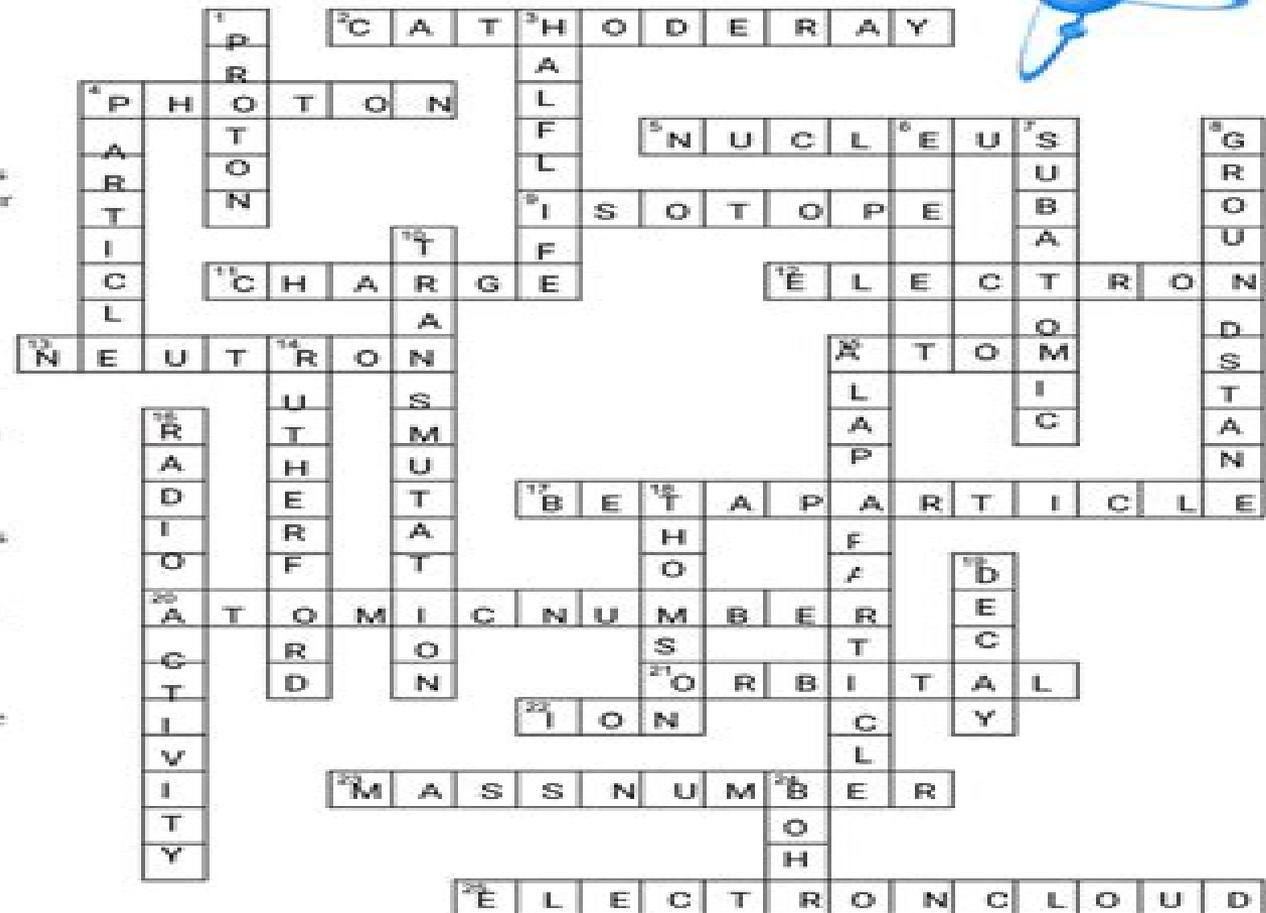


Atomic Structure



ACROSS

2. A beam of electrons.
4. The fundamental particle of visible light.
5. The center of an atom.
9. The term given to atoms that have the same number of protons but different numbers of neutrons.
11. How positive or negative an atom is.
12. A negatively charged subatomic particle.
13. A subatomic particle with no charge.
15. The smallest unit of an element.
17. Electron emitted by radioactive decay.
20. The number of protons in an atom.
21. The region of space around the nucleus where an electron is likely to be found.
22. An atom that has a positive or negative charge due to the gain or loss of one or more electrons.
23. The total number of protons and neutrons in the nucleus of an atom.
25. Empty space around the nucleus of an atom where electrons travel.



DOWN

1. A subatomic particle with a positive charge.
3. The time for a radioactive substance to lose half of its radioactivity.
4. A minute quantity of matter.
6. A form of matter that cannot be chemically broken down into

simpler substances.

7. Smaller than or occurring within an atom.
8. The lowest possible energy state of an atom.
10. Changing of one element to another through radioactive decay.
14. New Zealand physicist who figured out subatomic structure.
15. Two neutrons and two protons

bound together and produced by radioactive decay.

16. Emission of particles from an unstable atom.
18. English physicist who discovered the electron.
19. To break into smaller pieces.
24. The Danish physicist who made essential contributions to the understanding of atomic structure.

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J Rink



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