

# 5

## ATOMIC STRUCTURE AND THE PERIODIC TABLE

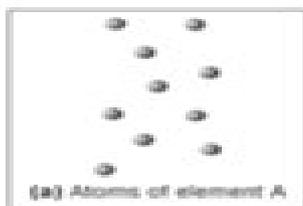
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### SECTION 5.1 ATOMS (pages 107–108)

*This section describes early atomic theories of matter and provides ways to understand the tiny size of individual atoms.*

#### ► Early Models of the Atom (pages 107–108)

- Democritus of Abdera, who lived in Greece during the fourth century B.C., suggested that matter is made up of tiny particles that cannot be divided. He called these particles atoms.
- List two reasons why the ideas of Democritus were not useful in a scientific sense. They did not explain chemical behavior, and they lacked experimental support because scientific testing was unknown at the time.
- The modern process of discovery about atoms began with the theories of an English schoolteacher named John Dalton.
- Circle the letter of each sentence that is true about Dalton's atomic theory.
  - All elements are composed of tiny, indivisible particles called atoms.
  - An element is composed of several types of atoms.
  - Atoms of different elements can physically mix together, or can chemically combine in simple, whole-number ratios to form compounds.
  - Chemical reactions occur when atoms are separated, joined, or rearranged; however, atoms of one element are never changed into atoms of another element by a chemical reaction.
- In the diagram, use the labels *mixture* and *compound* to identify the mixture of elements A and B and the compound that forms when the atoms of elements A and B combine chemically.



# Atomic Structure Periodic Table Answer Key

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