

Date Performed: 10/26/09

Date Submitted: 11/02/09

Instructor:

Name:

Partner:

Acid-Base Titrations

Objective

- To standardize a sodium hydroxide solution (NaOH) with a hydrochloric acid solution (HCl) of known concentration by the reaction.
- To determine the molar mass of an unknown, monoprotic acid
- To determine the percentage of Acetylsalicylic Acid in a tablet of aspirin.

Experimental Data

Part A-Standardization of Sodium Hydroxide Solution

Titration	Vol NaOH (mL)
1	20.00
2	20.21

Part B-Molar Mass of an Unknown Monoprotic Acid

Sample	Mass Beaker (g)	Mass Sample + Beaker (g)	Vol NaOH (mL)
1	119.50	120.50	20.95
2	88.32	89.32	19.95

Part C-Percentage of Acetylsalicylic Acid in a Commercial Aspirin Tablet

Mass Beaker (g)	Mass Tablet +Beaker (g)	Vol NaOH (mL)
122.08	122.46	8.49

Sample Calculations

1) Molarity NaOH

$(\text{Molarity HCl})(\text{Volume HCl}) = (\text{Molarity NaOH})(\text{Volume NaOH})$

$(0.200\text{M})(25.00\text{mL}) = (\text{Molarity NaOH})(20.00\text{mL})$

$5.00 = X(20.00)$

Trail 1: Molarity of NaOH = 0.250M NaOH

Trail 2: Molarity of NaOH = 0.247M NaOH

Average = 0.249M NaOH

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AW Chickering



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Integrated Approach to Coordination Chemistry Rosemary A. Marusak, Kate Doan, Scott D. Cummings, 2007-03-07
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