

B. Titration of the Acid of unknown concentration with the Standardized NaOH

1. Record your unknown ID.
2. Obtain less than 10mL of an unknown acid in a clean, dry 50-mL beaker.
3. Rinse the 10.00-mL pipet with three 2-mL portions of your unknown solution. Obtain ~ 40mL of unknown acid and pipet 10.00mL of it into the rinsed 250-mL flask.
4. Repeat the titration procedure above using the acid of unknown concentration in place of the standard acid.

Set up a data table to organize your data for the H₂SO₄/NaOH titration.

	Trial 1	Trial 2	Trial 3
Unknown ID			
V _{H₂SO₄} (mL)	10.00	10.00	10.00
Avg M _{NaOH} (M), from Part A			
V _{initial, NaOH} (mL)			
V _{final, NaOH} (mL)			
Δ V _{NaOH} (mL)			
M _{H₂SO₄} (M)			

Calculations

In these titrations, one mole of HCl reacts with one mole of NaOH, and one mole of H₂SO₄ reacts with two moles of NaOH. The molarity, M, of the acid or the base is expressed in moles per liter.

In part A the molarity of the HCl solution is known. The molarity of NaOH may be calculated from the results of part A through the following stoichiometry reflecting the overall equation:



1. M_{NaOH}, the molarity of the NaOH calculated from the data. Show one calculation of the NaOH molarity in detail
2. Average value of the three determinations of NaOH molarity.
3. Percent range of the three determinations:

$$\% \text{ range} = \frac{|\text{largest value} - \text{smallest value}|}{\text{average value}} \times 100\%$$

In part B the molarity of the NaOH solution determined from part A is used to compute the molar concentration of the H₂SO₄ in the solution of unknown concentration:



1. M_{H₂SO₄}, the molarity of the H₂SO₄ calculated from the data. Show one calculation of the H₂SO₄ molarity in detail
2. Average value of the three determinations of H₂SO₄ molarity.
3. Percent range of the three determinations:

$$\% \text{ range} = \frac{|\text{largest value} - \text{smallest value}|}{\text{average value}} \times 100\%$$

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Energy Research Abstracts ,1980 **The Software Encyclopedia** ,1988 Current Chemical Papers Chemical Society (Great Britain),1964 A classified world list of new papers in pure chemistry **The Effects of Laboratory Curriculum and Instruction on Undergraduate Students' Understanding of Chemistry** Dawn Kerry Rickey,1999 School Library Journal ,1985

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Table of Contents Acid Base Titrations Pre Lab Answers

1. Understanding the eBook Acid Base Titrations Pre Lab Answers
 - The Rise of Digital Reading Acid Base Titrations Pre Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Base Titrations Pre Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Base Titrations Pre Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Base Titrations Pre Lab Answers
 - Personalized Recommendations

- Acid Base Titrations Pre Lab Answers User Reviews and Ratings
- Acid Base Titrations Pre Lab Answers and Bestseller Lists
- 5. Accessing Acid Base Titrations Pre Lab Answers Free and Paid eBooks
 - Acid Base Titrations Pre Lab Answers Public Domain eBooks
 - Acid Base Titrations Pre Lab Answers eBook Subscription Services
 - Acid Base Titrations Pre Lab Answers Budget-Friendly Options
- 6. Navigating Acid Base Titrations Pre Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Acid Base Titrations Pre Lab Answers Compatibility with Devices
 - Acid Base Titrations Pre Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid Base Titrations Pre Lab Answers
 - Highlighting and Note-Taking Acid Base Titrations Pre Lab Answers
 - Interactive Elements Acid Base Titrations Pre Lab Answers
- 8. Staying Engaged with Acid Base Titrations Pre Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Base Titrations Pre Lab Answers
- 9. Balancing eBooks and Physical Books Acid Base Titrations Pre Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Base Titrations Pre Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acid Base Titrations Pre Lab Answers
 - Setting Reading Goals Acid Base Titrations Pre Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acid Base Titrations Pre Lab Answers
 - Fact-Checking eBook Content of Acid Base Titrations Pre Lab Answers

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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