

PRE-LAB ASSIGNMENT

1. Determine the mass of $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ needed to make 250 mL of 0.080 M $\text{H}_2\text{C}_2\text{O}_4$.

2. Determine the volume of 6.0 M NaOH needed to make 500 mL of 0.10 M NaOH.

3. If 32.96 mL of a NaOH solution are required to titrate 25.00 mL of 0.08121 M $\text{H}_2\text{C}_2\text{O}_4$ to the end point, what is the molarity of the sodium hydroxide solution?

4. If 24.92 mL of the NaOH solution from Question 3 are required to titrate 0.1932 g of an unknown diprotic ($n = 2$) acid, what is the molar mass of the acid?

5. In the standardization of the NaOH solution, why is it important to rinse the wet pipet with small portions of the standard oxalic acid solution?

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Hygienic Laboratory (U.S.)



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