

17 • Acid-Base Equilibria

PRACTICE TEST

- What is the $[H^+]$ when $[OH^-] = 8.1 \times 10^{-9}$?
 a) $8.1 \times 10^{-5} M$ d) $3.6 \times 10^{-8} M$
 b) $1.0 \times 10^{-7} M$ e) $8.1 \times 10^{-2} M$
 c) $1.2 \times 10^{-6} M$ $[H^+] = \frac{K_w}{[OH^-]} = \frac{1 \times 10^{-14}}{8.1 \times 10^{-9}} = 1.2 \times 10^{-6}$
- What is the $[H^+]$ when $[OH^-] = 3.3 \times 10^{-9}$?
 a) $3.0 \times 10^{-6} M$ d) $6.6 \times 10^{-5} M$
 b) $1.0 \times 10^{-7} M$ e) $3.3 \times 10^{-9} M$
 c) $3.3 \times 10^{-5} M$ $[H^+] = \frac{K_w}{[OH^-]} = \frac{1 \times 10^{-14}}{3.3 \times 10^{-9}} = 3.3 \times 10^{-6}$
- What is the $[H^+]$ in a 0.0025 M HCl solution?
 a) $1.0 \times 10^{-3} M$ d) $3.6 \times 10^{-5} M$
 b) $4.0 \times 10^{-12} M$ e) need more info
 c) $2.5 \times 10^{-3} M$ $HCl \rightarrow H^+ + Cl^-$
 $[H^+] = .0025 = 2.5 \times 10^{-3}$
- What is the $[OH^-]$ in a 0.0050 M HCl solution?
 a) $5.0 \times 10^{-3} M$ d) $6.6 \times 10^{-5} M$
 b) 1.0 M e) $2.0 \times 10^{-12} M$
 c) $1.0 \times 10^{-7} M$ $[H^+] = .0050$
 $[OH^-] = \frac{K_w}{[H^+]} = \frac{1 \times 10^{-14}}{.0050} = 2.0 \times 10^{-12}$
- A solution in which $[H^+] = 10^{-8}$ has a pH of 8 and is basic.
 a) 8, acidic d) -8, neutral
 b) 6, basic e) 8, basic
 c) -6, basic
- What is the pH of a 0.00030 M HNO_3 solution?
 a) 8.11 d) 4.48
 b) 3.00 e) none of these
 c) 3.52 $[H^+] = 3 \times 10^{-4}$
 $pH = -\log [H^+] = 3.52$
- What is the pH of a 0.0060 M KOH solution?
 a) 5.12 d) 8.88
 b) 2.22 e) 7.00
 c) 11.78 $[OH^-] = 6.0 \times 10^{-3}$
 $pOH = 2.22$
 $pH = 14 - pOH = 11.78$

- A sample of lemon juice is found to have a pH of 2.55. What is the H^+ concentration of the juice?
 $[H^+] = 10^{-pH} = 10^{-2.55}$
 a) 0.0035 M d) 0.0080 M
 b) 0.0028 M e) 355 M
 c) 11.6 M
- A sample of milk is found to have a pH of 6.60. What is the OH^- concentration of the milk?
 a) $2.5 \times 10^{-22} M$ d) $4.0 \times 10^{-8} M$
 b) $1.0 \times 10^{-7} M$ e) $2.5 \times 10^{-7} M$
 c) $5.0 \times 10^{-7} M$ $[H^+] = 10^{-6.60} = 2.5 \times 10^{-7}$
 $[OH^-] = \frac{K_w}{[H^+]}$
- What is the concentration of OCl^- in a 0.60 M solution of HOCl? $K_a = 3.1 \times 10^{-8}$.
 a) $1.8 \times 10^{-4} M$ d) $1.4 \times 10^{-4} M$
 b) $7.1 \times 10^{-11} M$ e) $1.1 \times 10^{-4} M$
 c) 0.40 M $[OCl^-] = [H^+] \cdot \frac{K_a \cdot .6}{x^2}$
~~see scratch paper~~
- What is the pH of a 0.020 M solution of hydrosulfuric acid, a diprotic acid?
 $K_{a1} = 1.1 \times 10^{-7}$ $K_{a2} = 1.0 \times 10^{-14}$ *only 1st dissociation is signif.*
 a) 7.00 d) 4.33 \therefore use K_{a1} only
 b) 9.67 e) 3.05
 c) 7.84
~~see scratch paper~~
- What is the concentration of CO_3^{2-} in a 0.010 M solution of carbonic acid? The relevant equilibria are:
 $H_2CO_3 \rightleftharpoons H^+ + HCO_3^-$ $K_{a1} = 4.3 \times 10^{-7}$
 $HCO_3^- \rightleftharpoons H^+ + CO_3^{2-}$ $K_{a2} = 5.6 \times 10^{-11}$
 a) $6.6 \times 10^{-5} M$ d) $7.5 \times 10^{-7} M$
 b) $5.6 \times 10^{-11} M$ e) $7.9 \times 10^{-7} M$
 c) $6.7 \times 10^{-11} M$
~~From the 1st dissociation, $[H^+] = [HCO_3^-]$
 the 2nd dissociation doesn't change this much.
 $K_{a2} = \frac{[H^+][CO_3^{2-}]}{[HCO_3^-]}$ but $[H^+] = [HCO_3^-]$ cancel!
 so $K_{a2} = [CO_3^{2-}] = 5.6 \times 10^{-11}$~~

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