

25. The number of  $\sigma$  bonds in  $N \equiv N$  is
- 1
  - 2
  - 3
  - 4
26. The elements in an ionic compound are held together by
- electrostatic forces of attraction.
  - van der Waals forces
  - the spin of paired electrons.
  - the formation of hybrid orbitals.
  - an electron pair.
27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode
- at which oxidation occurs.
  - which attracts cations.
  - at which electrons are supplied to the solution.
  - at which reduction occurs.
28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?
- 47
  - 62
  - 93
  - 186
29. In a galvanic (voltaic) cell in which the reaction is  $Cd + Cu^{2+} \rightarrow Cu + Cd^{2+}$  and the ions are at unit concentration (activity), the cell potential is
- $$Cd \rightarrow Cd^{2+} + 2e^- \quad 0.4021 \text{ V}$$
- $$Cu \rightarrow Cu^{2+} + 2e^- \quad -0.344 \text{ V}$$
- 0.1383 V
  - 0.4021 V
  - 0.344 V
  - 0.7461 V
  - 0.3677 V
30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?
- $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$
  - $H_2(g) + Cl_2(g) \rightleftharpoons 2HCl(g)$
  - $2NO(g) + O_2(g) \rightleftharpoons 2NO_2(g)$
  - $COCl_2(g) \rightleftharpoons CO(g) + Cl_2(g)$

Standard Potentials	$E^\circ$
$Mg \rightarrow Mg^{2+} + 2e^-$	2.37V
$Al \rightarrow Al^{3+} + 3e^-$	1.66V
$Zn \rightarrow Zn^{2+} + 2e^-$	0.76V
$Fe \rightarrow Fe^{2+} + 2e^-$	0.44V
$Cu \rightarrow Cu^{2+} + 2e^-$	0.34V
$Ag \rightarrow Ag^+ + e^-$	-0.80V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.
- Mg and Ag
  - Mg and Fe
  - Zn and Cu
  - Al and Ag
  - Mg and Al
32.  $E = E^\circ - 0.059/n \log Q$  (Nernst equation)  
 $[H^+] = 1.0 \text{ M}$  initially,  $P_{O_2} = 1.0 \text{ atm}$
- $$4e^- + O_2(g) + 4H^+(aq) \rightleftharpoons 2H_2O(l) \quad E^\circ = 1.23 \text{ V}$$
- Based on the information above, which statement is correct?
- $n = 1$ , since one mole of oxygen is being considered.
  - Addition of base should result in an  $E$  value, which is less than 1.23 V.
  - $E$  is independent of the pH of the solution.
  - $Q = \frac{[H_2O]^2}{[O_2][H^+]^4}$
33. The equilibrium constant for the gaseous reaction  $C + D \rightleftharpoons E + 2F$  is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will
- proceed at equal rates in both directions.
  - proceed more rapidly to form E and F.
  - proceed more rapidly to form C and D.
  - not occur in either direction.

Compound	$\Delta G_f^\circ$ , kJ/mol
$H_2O(l)$	-237
$H_2O(g)$	-229

34. At 298 K the equilibrium constant for  $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(l)$
- is larger than the  $K_{eq}$  for  $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(g)$
  - will have a value of 1.0 at equilibrium.
  - cannot be computed since data on  $O_2$  and

# Acs Standardized Exam General Chemistry Ii

**Michal Rosen-Zvi**



**Acs Standardized Exam General Chemistry Ii:**

This book delves into Acs Standardized Exam General Chemistry Ii. Acs Standardized Exam General Chemistry Ii is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Acs Standardized Exam General Chemistry Ii, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Acs Standardized Exam General Chemistry Ii
- Chapter 2: Essential Elements of Acs Standardized Exam General Chemistry Ii
- Chapter 3: Acs Standardized Exam General Chemistry Ii in Everyday Life
- Chapter 4: Acs Standardized Exam General Chemistry Ii in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Acs Standardized Exam General Chemistry Ii. The first chapter will explore what Acs Standardized Exam General Chemistry Ii is, why Acs Standardized Exam General Chemistry Ii is vital, and how to effectively learn about Acs Standardized Exam General Chemistry Ii.
3. In chapter 2, the author will delve into the foundational concepts of Acs Standardized Exam General Chemistry Ii. This chapter will elucidate the essential principles that need to be understood to grasp Acs Standardized Exam General Chemistry Ii in its entirety.
4. In chapter 3, the author will examine the practical applications of Acs Standardized Exam General Chemistry Ii in daily life. This chapter will showcase real-world examples of how Acs Standardized Exam General Chemistry Ii can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Acs Standardized Exam General Chemistry Ii in specific contexts. This chapter will explore how Acs Standardized Exam General Chemistry Ii is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Acs Standardized Exam General Chemistry Ii. This chapter will summarize the key points that have been discussed throughout the book.
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