

H. 59. Alpha particles with their high ionizing charge but low penetrating ability can best be shielded by?

- a. paper      b. clothing      c. water      d. lead

Alpha  $\alpha$  particles - paper is a sufficient shield

Beta  $\beta$  particles - heavy clothing is sufficient to shield

Gamma rays - lead is required to shield



D. 60. The half-life of Se-121 is 10 days. If you started with 40 grams of this isotope, how much would you have left after 30 days?

- a. 10 g      b. none      c. 15 g      d. 5 g

Mass of Se-121	Number of half-lives	Time
40 g	0	0
20 g	1	10 days
10 g	2	20 days
5 g	3	30 days

$\oplus \uparrow$  = particles go faster  
 $\oplus \downarrow$  = particles go slower

**PERIODIC TRENDS**

ATOMIC RADIUS: decreased  $\rightarrow$   
 increased  $\downarrow$

METALLIC CHARACTER:

decreases  $\downarrow$   
 increases  $\rightarrow$

IONIZATION  $\oplus$  decreases  $\downarrow$   
 increases  $\rightarrow$



Isotopes - atoms with same # of protons, but different # of neutrons.

EXAMPLE: Ne-20  $\rightarrow$   $^{20}_{10}\text{Ne}$   
 Ne-21  $\rightarrow$   $^{21}_{10}\text{Ne}$   
 Ne-22  $\rightarrow$   $^{22}_{10}\text{Ne}$

Dilution Equation  $M_1 V_1 = M_2 V_2$

• mass % =  $\frac{\text{mass solute}}{\text{mass of solution (solute + solvent)}} \times 100\%$

• Molarity (M) =  $\frac{\text{mol solute}}{\text{L solution}}$

• molality (m) =  $\frac{\text{mol solute}}{\text{kg solvent}}$

- #protons - #electrons = Charge #
- #neutrons = A (mass #) - Z (atomic #)
- if no charge, the #protons + #electrons = same
- if + (positive) charge (cation) = #electrons is less than #protons
- if - (negative) charge (anion) = #electrons is greater than #protons
- Protons  $\neq$  Charge  $\rightarrow$  they give identity to the element.

# Acs General Chemistry Exam Review

**Sabine Zange**



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