

Fisicoquímica.

Problemario 1/3

- For steam at 500°C and 10 MPa, using the Mollier diagram,
 - Compute the Joule-Thomson coefficient $\mu_{JT} = \left(\frac{\partial T}{\partial p}\right)_H$
 - Compute the coefficient $\kappa_s = \left(\frac{\partial T}{\partial p}\right)_s$
- Evaluate the difference $\left(\frac{\partial v}{\partial T}\right)_p - \left(\frac{\partial v}{\partial T}\right)_v$ for the ideal and van der Waals gases, and for a gas that obeys the virial equation of state.
- One hundred cubic meters of carbon dioxide initially at 150°C and 50 bar is to be isothermally compressed in a frictionless piston-and-cylinder device to a final pressure of 300 bar. Calculate
 - The volume of the compressed gas
 - The work done to compress the gas
 - The heat flow on compressionassuming carbon dioxide a. is an ideal gas, b. Obeys the principle of corresponding states, c. Obeys the Peng-Robinson equation of state.
- Ethylene at 30 bar and 100°C passes through a heater-expander and emerges at 20 bar and 150°C. There is no flow of work into or out of the heater-expander, but heat is supplied. Assuming that ethylene obeys the Peng-Robinson equation of state, compute the flow of heat into the heater-expander per mole of ethylene.



- Calculate ΔA and ΔG when a mole of water vapor initially at 200°C and 1 bar undergoes a cyclic process for which $W = 145$ J.
- a) Use the equations $C_p = \left(\frac{\partial H}{\partial T}\right)_p$, $\frac{\partial}{\partial T} \left(\frac{\partial H}{\partial p}\right) = \frac{\partial}{\partial p} \left(\frac{\partial H}{\partial T}\right)$, and $\left(\frac{\partial H}{\partial p}\right)_T = -T \left(\frac{\partial v}{\partial T}\right)_p + V = -TV\alpha + V$ to show that $\left(\frac{\partial C_p}{\partial p}\right)_T = -T \left(\frac{\partial^2 v}{\partial T^2}\right)_p$. The volumes of substances increase approximately linearly with T, so $\frac{\partial^2 v}{\partial T^2}$ is usually quite small. Consequently, the pressure dependence of C_p can usually be neglected unless one is dealing with high pressures.

The volume of Hg in the temperature range 0°C to 100°C at 1 atm is given by $V = V_0(1 + \alpha t + bt^2)$, where $\alpha = 0.18182 \times 10^{-3} \text{ } ^\circ\text{C}^{-1}$, $b = 0.78 \times 10^{-8} \text{ } ^\circ\text{C}^{-2}$, V_0 is the volume at 0°C, and t is the Celsius temperature. The density of mercury at 1 atm and 0°C is 13.596 g/cm³. b) Calculate $\left(\frac{\partial C_p}{\partial p}\right)_T$ for Hg at 25°C and 1 atm. c) Given $C_{p,m} = 6.66 \text{ cal mol}^{-1} \text{ K}^{-1}$ for Hg at 1 atm and 25°C, estimate $C_{p,m}$ of Hg at 25°C and 10⁴ atm.
- Calculate ΔG for the isothermal compression of 30.0 g of water from 1.0 atm to 100.0 atm at 25°C; neglect the variation of v with p .
- Suppose 1.00 mol of water initially at 27°C and 1 atm undergoes a process whose final state is 100°C and 50 atm. Calculate ΔH , ΔU , and ΔS considering that the temperature and pressure variations of α , κ , and C_p can be neglected.

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