

PHYSICAL PHARMACY

PHYSICAL CHEMISTRY – branch of chemistry concerned with physical properties and structure of matter and with laws and theories of physical and chemical changes

PHYSICAL PHARMACY – application of physical chemistry in pharmacy; study of physicochemical properties of substances used in drug formulation

I. FORCES OF ATTRACTION

A. INTRAMOLECULAR – intra (within) the molecule

- **Ionic Bond** – transfer of electrons; metal (electron donor) + non-metal (electron acceptor); (ex. Na: $p^+ = 11$; $e^- = 11 \rightarrow p^+ = 11$; $e^- = 10 \rightarrow$ charge = +1)

○ PERIODIC TRENDS REPRESENTING BONDS

- **Ionization Energy** – energy required remove electron (low M; high NM)
- **Electronegativity** – ability to attract electrons (low M; high NM)

- **Covalent Bonds** – sharing of electrons; non-metal + non-metal; polarity – difference in electronegativity (symmetrical = non-polar)

- **Polar** – unequal sharing of electrons; dipole moment – (ex. $H^{\delta+}-Br^{\delta-}$)
- **Non Polar** – equal sharing of electrons (ex. $Br-Br$)

B. INTERMOLECULAR – inter (between) the molecules

- **Van der Waals**

- **Dipole-Dipole** $\rightarrow P + P$; Keesom Forces; orientation effect
- **Dipole-Induced Dipole** $\rightarrow P + NP$; Debye Forces; inductive effect
- **Induced Dipole-Induced Dipole** $\rightarrow NP + NP$; London Forces; dispersive effect

- **Ion-Induced Dipole** $\rightarrow ion + NP$

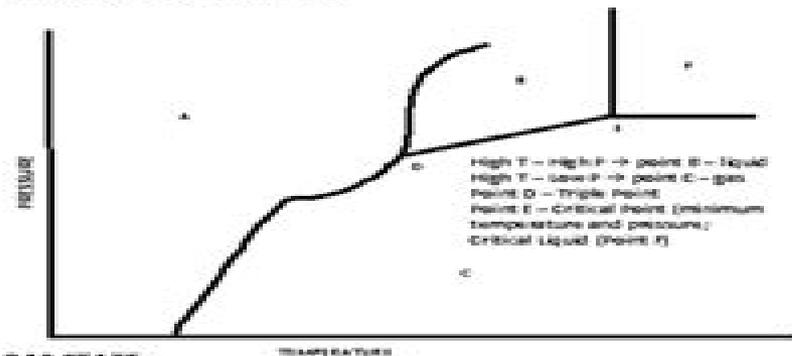
$I_2 \rightarrow NP$ (solubility = 1:2960)

$I_2 + NaI/KI$ (solubilizing agent) $\rightarrow I_3 + I_2 =$ increase solubility

- **Hydrogen Bond** – hydrogen atom + electronegative atom (N, O, F)



PHASE DIAGRAM



GAS STATE:

- **Kinetic Molecular Theory**

- Gases are composed of small particles called atoms; total volume is negligible in relation to space where they are confined
- Particles do not attract one another but move with complete independence
- Particles exhibit continuous random motion owing to their kinetic energy
- Gases exhibit perfect elasticity (no energy exchange)

II. GAS LAWS

- at standard temperature (Kelvin = $^{\circ}C + 273$) and pressure (1 atm = 760 mmHg or torr; 760 mmHg; 1.01325×10^5 dynes/cm²; 1.01325×10^5 Pa or N/m²; 0.1325 bar)

- **Boyle's Law** – constant temperature; increase P = decrease V; inversely related

$$P \propto \frac{1}{V}$$

- * density = increase altitude; decrease volume; $P_1V_1 = P_2V_2$

- **Charles' Law** – constant pressure; increase T = increase V

$$T \propto V; \frac{T_1}{T_2} = \frac{V_1}{V_2} \text{ or } \frac{T_1}{V_1} = \frac{T_2}{V_2}$$

- * H_2O – 1.0g/ml $\rightarrow 4^{\circ}C$

- **Gay Lussac's Law** \rightarrow constant volume

$$P \propto T; \frac{P_1}{P_2} = \frac{T_1}{T_2} \text{ or } \frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- * $121^{\circ}C$ – 15-20psi

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DJ Losen



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