

Chapter Test B

Atomic Physics

MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. What is the frequency of a photon with an energy of $1.00 \times 10^{-19} \text{ J}$?
 ($h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$)
- | | |
|-------------------------------------|-------------------------------------|
| a. $1.00 \times 10^{21} \text{ Hz}$ | c. $3.00 \times 10^{12} \text{ Hz}$ |
| b. $2.00 \times 10^{21} \text{ Hz}$ | d. $4.00 \times 10^{12} \text{ Hz}$ |
- _____ 2. Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
- | | |
|----------|----------|
| a. 1 : 1 | c. 3 : 1 |
| b. 2 : 1 | d. 4 : 1 |
- _____ 3. A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
- | | |
|-----------|-----------|
| a. 4.6 eV | c. 1.4 eV |
| b. 4.8 eV | d. 2.4 eV |
- _____ 4. What causes the bright lines in the emission spectrum of an element to occur?
- Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
- _____ 5. What causes the dark lines in the absorption spectrum of an element to occur?
- Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

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