

## Thermochemistry Quiz 1

Questions 1-4 (1 pt. each); Questions 5-6 (2 pts. Each)

- \_\_\_\_\_ 1. An endothermic reaction causes the surroundings to
- warm up.
  - become acidic.
  - condense.
  - decrease in temperature.
  - release CO<sub>2</sub>.
- \_\_\_\_\_ 2. According to the first law of thermodynamics:
- energy is neither lost nor gained in any energy transformations.
  - perpetual motion is possible.
  - energy is conserved in quality but not in quantity.
  - energy is being created as time passes. We have more energy in this universe now than when time began.
- \_\_\_\_\_ 3. A calorimeter
- is equal to the molar enthalpy of a reaction.
  - is a dieting aid.
  - is an indicator of a spontaneous reaction.
  - is a device used to measure the transfer of heat energy.
  - is useful in measuring the amount of heat released by endothermic reactions.
- \_\_\_\_\_ 4. The specific heat of gold is 0.129 J/g•°C. What is the molar heat capacity of gold?
- 0.039 J/mol•°C
  - 0.129 J/mol•°C
  - 25.4 J/mol•°C
  - 39.0 kJ/mol•°C
  - 197 J/mol•°C
- \_\_\_\_\_ 5. Copper metal has a specific heat of 0.385 J/g•°C. Calculate the amount of heat required to raise the temperature of 22.8 g of Cu from 20.0°C to 875°C.
- $1.97 \times 10^{-2}$  J
  - $1.0 \times 10^{-2}$  J
  - 329 J
  - 7.51 kJ
  - 10.5 kJ
- \_\_\_\_\_ 6. The combustion of butane produces according to the equation
- $$2\text{C}_4\text{H}_{10}(\text{g}) + 13\text{O}_2(\text{g}) \rightarrow 8\text{CO}_2(\text{g}) + 10\text{H}_2\text{O}(\text{l}) \quad \Delta H^\circ_{\text{rxn}} = -5,314 \text{ kJ}$$
- How many grams of butane must be burned to release 6,375 kJ of heat?
- 1.20 g
  - 139 g
  - 0.0413 g
  - 69.7 g
  - 97.8 g

Questions 7-8: Label the following processes as either *exothermic* or *endothermic*. (1 pt. each)

7. Dry ice (solid CO<sub>2</sub>) sits at room temperature and sublimates to CO<sub>2</sub> gas. \_\_\_\_\_
8. A block of solid aluminum is heated until it melts. \_\_\_\_\_

# Thermochemistry Practice Test A Answers

**David Wilson**



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