

DATA AND CALCULATIONS

A. pH of Acid Solutions:

1. Strong Acid	[HCl] ₁	[HCl] ₂
Measured pH	<u>1.35</u>	<u>1.95</u>
Molarity	_____	_____
Calculated pH	_____	_____
2. Weak Acid	[HC ₂ H ₃ O ₂] ₁	[HC ₂ H ₃ O ₂] ₂
Measured pH	<u>2.84</u>	<u>3.20</u>
Molarity	_____	_____
Calculated pH	_____	_____

Compare and explain the observed values of pH for these two acids.

Which acid pH changed the most on dilution? Explain why (consider the equilibrium occurring).

Ph Properties Of Buffer Solutions Lab Calculations

JG Myers



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