

THERMODYNAMICS

JAM-2010

(1) The molar internal energy of a gas at temperature T is $U_m(T)$. The molar internal energy at $T=0$ is $U_m(0)$. The correct expression that relates these two with appropriate contribution is ---

(A) $U_m(T) = U_m(0) + 3RT$ [Linear molecule, translation only]

(B) $U_m(T) = U_m(0) + \frac{7}{2}RT$ [Linear molecule, translation and rotation only]

(B) $U_m(T) = U_m(0) + \frac{5}{2}RT$ [non linear molecule, translation and rotation only]

(D) $U_m(T) = U_m(0) + RT$ [non linear molecule, translation only]

Ans: For Linear molecule translational D.O.F =3, rotational D.O.F = 2

For each translational D.O.F Energy = $\frac{1}{2}NkT$ so, total translational Energy = $\frac{3}{2}NkT$

For each rotational D.O.F Energy = NkT so, total rotational Energy = $2NkT$

So total translational and rotational contribution is = $\frac{7}{2}NkT = \frac{7}{2}RT$

So the option (B) is correct.

(2) The work done during the free expansion of one mole of an ideal gas at 27°C to twice its original volume is (given : 2494 J mol^{-1} , $\ln 2 = 0.7$, $\log 2 = 0.3$)

(A) 1746 J mol^{-1}

(B) -1746 J mol^{-1}

(C) Zero

(D) 748.2 J mol^{-1}

Ans: For free expansion work done by the system is zero

So option (C) is the correct answer

(3) Calculate ΔH_m and ΔS_m for the process



Assuming that at 273 K the molar enthalpy of fusion of ice is 6006 J mol^{-1} , the heat capacity $C_{p,m}(s)$ of ice is $38 \text{ JK}^{-1}\text{mole}^{-1}$ and the heat capacity $C_{p,m}(l)$ of liquid water is $76 \text{ JK}^{-1}\text{mole}^{-1}$. Consider the heat capacity to be constants. (given $\ln 263 = 5.57$ and $\ln 273 = 5.61$)

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RD Boyd



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