

## LAB 06 – DOUBLE DISPLACEMENT REACTIONS

### Discussion - Double Replacement Reactions (Refer to class notes from Chapter 7)

All reactions in this part of the experiment are double replacement reactions.

Double replacement reactions, sometimes called "double displacement", "exchange reactions" or "trading partners" are very common. Generally, they occur when ionic compounds are dissolved in aqueous (water) solution. They have the general form:



where A and B are positive ions and X and Y are negative ions.

A reaction will take place if one of the following is formed in the reaction:

- a **precipitate** (what you see is a solid material that clouds the solution).
- a **gas** (you see lots of bubbles coming out of solution).
- water** (You will feel the test tube get warm as heat is given off by the reaction).

These are known as the **Driving Force** of the reaction.

If the reaction produces neither a precipitate, nor a gas, nor a molecular substance, then **no reaction** has taken place. If a reaction does take place, then moving from a balanced molecular equation to a net ionic equation allows us to determine exactly what the chemistry was in the reaction, eliminating the non-reacting (spectator) ions and/or molecules.

The following double displacement reactions were carried out in lab and the listed observations were made.

Reagents	Observations
1. $\text{NaCl}_{(aq)} + \text{AgNO}_{3(aq)}$	<i>White precipitate formed</i>
2. $\text{Na}_2\text{CO}_{3(aq)} + \text{HCl}_{(aq)}$	<i>Bubbles form, solution got warm</i>
3. $\text{NaOH}_{(aq)} + \text{HCl}_{(aq)}$	<i>Solution got warm</i>
4. $\text{Na}_2\text{CO}_{3(aq)} + \text{CaCl}_{2(aq)}$	<i>White precipitate formed</i>
5. $\text{FeCl}_{3(aq)} + \text{NH}_4\text{OH}_{(aq)}$	<i>Reddish brown precipitate formed</i>
6. $\text{CuSO}_{4(aq)} + \text{Zn}(\text{NO}_3)_2(aq)$	<i>No Reaction</i>

# Double Displacement Reaction Lab Answers

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