

Double-Replacement Precipitation Lab

Introduction: Double-Replacement reactions occur when two ionic solutions mix to form an insoluble solid. If all cations and anions are soluble then no reaction will occur. During the lab procedure, several ionic solutions will be mixed to determine if a precipitate reaction will occur. Those reactions that do not form a precipitate will be marked for No Reaction (NR). Students will practice writing and balancing double-replacement reactions.

Pre-Lab questions:

1. How would you describe a double replacement reaction to a friend who is struggling with reaction types?
2. What is a precipitate?

Materials:

- 0.2M sodium bromide, NaBr
- 0.2M sodium carbonate, Na₂CO₃
- 0.2M sodium phosphate, Na₃PO₄
- 0.2M sodium sulfate, Na₂SO₄
- 0.2M sodium hydroxide, NaOH
- Transparency paper
- 0.2M lithium chloride, LiCl
- 0.2M copper(II) chloride, CuCl₂
- 0.2M zinc chloride, ZnCl₂
- 0.2M strontium chloride, SrCl₂
- 0.2M barium chloride BaCl₂
- Background paper Template

Safety: You MUST wear goggles during this lab activity

Procedure:

1. Place the overhead sheet over the background paper.
2. Using the Data Table as a guide, mix every possible pair of solutions in a separate square of the background paper. **Use 1 or 2 drops of each solution.** Do NOT contaminate the individual pipets with different solutions.
3. Mix the solutions with a rinsed toothpick if you are uncertain if a precipitate formed. Rinse the toothpick after EACH mixing – be careful of which end of the rod you use so that you do not touch the chemicals or contaminate your samples!
4. Observe each mixture carefully for signs of a precipitate. (Some precipitates are light in color and hard to see so you may want to slide the transparency over the black lab top on occasion.) Record your observations (presence or absence of a precipitate, and other characteristics such as color, texture, etc.) in the Data Table. If nothing happens, write NR for no reaction.
5. Rinse and wipe the transparency sheet into the sink. Dry the sheet.

Observations: Write the observations you made from combining these materials in the spaces in the table below. Note any color changes or precipitate (cloudiness/solid). If no change is observed write NR for no reaction.

	NaBr	Na ₂ CO ₃	Na ₃ PO ₄	Na ₂ SO ₄	NaOH
LiCl					
CuCl ₂					
ZnCl ₂					
SrCl ₂					
BaCl ₂					

Double Replacement Reaction Lab Conclusion Answers

**Eugene LeMay, Jr., Herbert Beall, Karen
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Brower, Prentice-Hall Staff**

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Safety-Scale Lab Exp Biochem 2e Spencer L. Seager, Michael R. Slabaugh, 1994-05 *Steps to Doing Science* Kristin Tuttle Bump, 2008 **Everything You Need to Ace Chemistry in One Big Fat Notebook** Workman Publishing, Jennifer Swanson, 2020-09-01 Chemistry No problem This Big Fat Notebook covers everything you need to know during a year of high school chemistry class breaking down one big bad subject into accessible units Learn to study better and get better grades using mnemonic devices definitions diagrams educational doodles and quizzes to recap it all Including Atoms elements compounds and mixtures The periodic table Quantum theory Bonding The mole Chemical reactions and calculations Gas laws Solubility pH scale Titrations Le Chatelier's principle and much more *Prentice Hall Chemistry*, 2000 **Novice Teacher Action** Anne Liu Kern, 2007 *Harcourt Science: Life science, units A and B*, 2002 *Ionic Liquids* Barbara Kirchner, 2009-12-01 Ionic liquids will never find application in industry I don't understand this fad for ionic liquids and there is no widespread interest in these systems are just three of quotes from the reports of referees for research proposals that I have received over the years I wonder what these people think today There are currently at least nine large scale industrial uses of ionic liquids including we now recognize the production of Caprolactam a monomer for the production of nylon 6 1 There has been a steady increase in the interest in ionic liquids for well over a decade and last year the number of papers and patents including ionic liquids was counted in the thousands This remarkable achievement has been built on the hard work and enthusiasm first of a small band of devotees but now of huge numbers of scientists all over the world who do not see themselves as specialists in ionic liquids The ionic liquids field continues to develop at an incredible rate No sooner do I think that I am on top of the literature than it turns out that a whole new area of work has emerged without me noticing Things that were once supposedly impossible in ionic liquids such as measuring the ¹H NMR of solutes are now widely applicable see Chapter 8 Hence collected volumes such as this are very welcome *Chemistry* Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02 *Harcourt Science: Physical science [grade] 6, units E and F, teacher's ed*, 2000 **A Constructivist Approach to the Study of Matter** Susan Hankins Fritzell, 2000 *Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science*, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction *Synthesis and Applications of Isotopically Labelled Compounds* 1994 J. Allen, R. Voges, 1995-08-15 This work provides a forum for the exchange of information between leading scientists and manufacturers involved in the synthesis and applications of compounds labelled with radioactive and stable isotopes Contains a general overview of the latest work in the field and its future significance **Ores and Metals**, 1906 *Chemical*

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