

Preparation of Soluble Salts

By Neutralization

25.0cm³ acid, as standard solution, is placed in conical flask using pipette. Add few drops of indicator & titrate with alkali from burette until indicator changes colour, showing all acid has just reacted. Volume of alkali added is measured. Prepare new 25.0cm³ acid again with pipette & add same volume of alkali as before to prevent excess alkali/acid because both reactant & product are aqueous. Next, the product is evaporated to dryness to obtain the salt.



1. Prepare a standard 25.0 cm³ of acid solution in a conical flask. Add a few drops of indicator. It is now about 10% of the volume of the acid.
2. The burette is now prepared. 25.0 cm³ of alkali solution is added to a burette. The burette is now about 10% of the volume of the acid. The burette is now about 10% of the volume of the acid.
3. The product solution is prepared by evaporating most of the water and crystallizing the salt.

By Reacting Metal with Acid



1. Add an excess of metal to acid slowly to a beaker. The metal reacts with the acid to form a salt solution and hydrogen gas. The metal is added in excess to ensure all the acid is used.



2. Filter the solution through a filter paper into a beaker. The residue is the metal that has reacted with the acid.



3. Heat the solution of your salt in an evaporating dish until most of the water has evaporated. The residue is the salt. The residue is the salt.



4. Allow the solution to cool and the crystals will form. The crystals are the salt. The crystals are the salt.

Only metals like zinc and magnesium, which moderately react with dilute acids, are used.

E.g. Reacting Zn with H₂SO₄ to prepare ZnSO₄



Zn is added to H₂SO₄ until in excess to ensure no more H₂SO₄ is present. Then the mixture is filtered off to separate Zn from ZnSO₄. The filtrate (ZnSO₄) is then placed in evaporating dish to evaporate most of water then it's cooled after ZnSO₄ crystals are formed. The crystals then filtered and squeezed between filter papers to dry.

By Reacting Insoluble Base with Acid

E.g. Reacting MgO with Acids



The same method as reaction of acid with metal is used, so refer to diagram and above explanation, substituting reactants and products.

By Reacting Carbonate with Acid

E.g. Reacting CaCO₃ with Acids



The same process is used as reaction of acid with metal, just that carbon dioxide is produced. Carbon dioxide can be tested by bubbling it into limewater which will turn limewater colourless to milky.

7.4 Properties and Uses of Ammonia

Ammonia and its Uses

Are produced from nitrogen reacted with hydrogen

For producing: fertilisers, nitric acid, nylon, dyes, cleaners and dry cell

The Manufacture of Ammonia: The Haber Process

The Process

Nitrogen and hydrogen are mixed together in ratio 1:3, where nitrogen is obtained from air and hydrogen is obtained from natural gas, and passed over iron-catalyst.



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