

# Modern Chemistry Chapter 15

## Test Review

- 25 multiple choice
  - definition of self-ionization
  - $[\text{H}_3\text{O}^+][\text{OH}^-] = 1 \times 10^{-14}$
  - definition and formula for calculating  $\text{pH} = -\log[\text{H}_3\text{O}^+]$
  - $\text{pH} > 7$  is a base;  $\text{pH} < 7$  is an acid;  $\text{pH} = 7$  is neutral
  - $\text{pH}$  range is normally 0 to 14
  - calculate  $\text{pH}$  from  $[\text{H}_3\text{O}^+]$
  - calculate  $[\text{H}_3\text{O}^+]$  from  $\text{pH}$
  - definitions of indicators, transition interval, titration, neutralization, endpoint, standard solution
  - acid-base titration monitors  $\text{pH}$
  - titration problems using  $V_a M_a \text{eq}_a = V_b M_b \text{eq}_b$

# Modern Chemistry Chapter 15 Mixed Review Answers

**M Planty**



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