

practice test ch 7 & 8

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. Which of the following best describes the rates of chemical reaction?
- most chemical reactions occur very slowly
 - most chemical reactions occur at moderate rates
 - most chemical reactions occur very rapidly
 - chemical reactions have a wide range of rates, from extremely fast to extremely slow
- _____ 2. Which of the following is true of the rates of most chemical reactions?
- the initial rate is cannot be measured
 - the initial rate is faster than the rate later in time
 - the initial rate is the same as the rate later in time
 - the initial rate is slower than the rate later in time
- _____ 3. Which of the following is true of effective collisions?
- the number of effective collisions decreases as the temperature is increased
 - the number of effective collisions determines the reaction rate
 - both a. and b.
 - Neither a. nor b.
- _____ 4. Which of the following is true of effective collisions?
- the number of effective collisions determines the reaction rate
 - the number of effective collisions increases as the temperature is increased
 - both a. and b.
 - Neither a. nor b.
- _____ 5. Which of the following is true?
- decreasing the activation increases the reaction rate
 - the activation energy for an exothermic reaction is negative
 - both a. and b.
 - Neither a. nor b.
- _____ 6. In an energy diagram for a chemical reaction what species appears at the highest energy?
- the catalyst
 - the products
 - the reactants
 - the transition state
- _____ 7. Given that the reaction $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$ is exothermic, which of the following is true of the reaction $2 \text{H}_2\text{O}(\text{g}) \rightarrow 2 \text{H}_2(\text{g}) + \text{O}_2(\text{g})$?
- its activation energy is lower than that of $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
 - its activation energy is the same as that of $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
 - its activation energy is higher than that of c. its activation energy is the same as that of $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
 - there is no relationship between its activation energy and that of $2 \text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2 \text{H}_2\text{O}(\text{g})$
- _____ 8. In an energy diagram for an endothermic chemical reaction which of the following is true?
- the energy of the products is lower than that of the reactants
 - the energy of the transition state is higher than that of the reactants
 - both a. and b.
 - Neither a. nor b.
- _____ 9. Consider a reaction such as $\text{A}_2(\text{g}) + \text{B}_2(\text{g}) \rightarrow 2\text{AB}(\text{g})$. What can we say about the mechanism of the reaction?
- if the reaction is fast it occurs in a single step
 - if the reaction is fast it occurs in multiple steps
 - if the reaction is slow it occurs in multiple steps

Modern Chemistry Chapter 7 Test Answer Key

**John O'M. Bockris, Amulya K.N.
Reddy, Maria E. Gamboa-Aldeco**

Modern Chemistry Chapter 7 Test Answer Key:

Elementary Modern Chemistry Wilhelm Ostwald, Harry Wheeler Morse, 1909 **Modern Electrochemistry 2B** John O'M. Bockris, Amulya K.N. Reddy, Maria E. Gamboa-Aldeco, 1998 This long awaited and thoroughly updated version of the classic text Plenum Press 1970 explains the subject of electrochemistry in clear straightforward language for undergraduates and mature scientists who want to understand solutions Like its predecessor the new text presents the electrochemistry of solutions at the molecular level The Second Edition takes full advantage of the advances in microscopy computing power and industrial applications in the quarter century since the publication of the First Edition Such new techniques include scanning tunneling microscopy which enables us to see atoms on electrodes and new computers capable of molecular dynamics calculations that are used in arriving at experimental values Chapter 10 starts with a detailed description of what happens when light strikes semi conductor electrodes and splits water thus providing in hydrogen a clean fuel There have of course been revolutionary advances here since the First Edition was written The book also discusses electrochemical methods that may provide the most economical path to many new syntheses for example the synthesis of the textile nylon The broad area of the breakdown of material in moist air and its electrochemistry is taken up in the substantial Chapter 12 Another exciting topic covered is the evolution of energy conversion and storage which lie at the cutting edge of clean automobile development Chapter 14 presents from a fresh perspective a discussion of electrochemical mechanisms in Biology and Chapter 15 shows how new electrochemical approaches may potentially alleviate many environmental problems *Lab Experiments Modern Chemistry* Harold Clark Metcalfe, 1982 **Modern inorganic chemistry** Joseph William Mellor, 1918

Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance **Introduction to Corrosion Science** Edward McCafferty, 2010-07-17 This textbook is intended for a one semester course in corrosion science at the graduate or advanced undergraduate level The approach is that of a physical chemist or materials scientist and the text is geared toward students of chemistry materials science and engineering This textbook should also be useful to practicing corrosion engineers or materials engineers who wish to enhance their understanding of the fundamental principles of corrosion science It is assumed that the student or reader does not have a

background in electrochemistry However the student or reader should have taken at least an undergraduate course in materials science or physical chemistry More material is presented in the textbook than can be covered in a one semester course so the book is intended for both the classroom and as a source book for further use This book grew out of classroom lectures which the author presented between 1982 and the present while a professorial lecturer at George Washington University Washington DC where he organized and taught a graduate course on Environmental Effects on Materials Additional material has been provided by over 30 years of experience in corrosion research largely at the Naval Research Laboratory Washington DC and also at the Bethlehem Steel Company Bethlehem PA and as a Robert A Welch Postdoctoral Fellow at the University of Texas The text emphasizes basic principles of corrosion science which underpin extensions to practice Laboratory Experiments to Accompany "Modern Chemistry," Fredus Nelson Peters,1902

Practical Chemistry Newton Henry Black,James Bryant Conant,1920 **Books in Print Supplement** ,2002 Modern Chemistry Nicholas D. Tzimopoulos,1990 **Handbook of Modern Chemistry** Charles Meymott Tidy,1878 **General Chemistry** Ralph H. Petrucci,1989 Modern Microcrystal Tests for Drugs Charles Clarke Fulton,1969 *Lab Experiments for Modern Chemistry* Tzimopoulo,1990 **Children's Books in Print** R R Bowker Publishing,Bowker,1999-12 **Modern Chemistry** H. Clark Metcalfe,Joseph F. Castka,1966 **Introduction to Modern Chemistry: Student Manual** Eugene Meyer,1979 **Essentials of Modern Chemistry** Charles Elwood Dull,1918

Handbook of Modern Pharmaceutical Analysis Satinder Ahuja,Stephen Scypinski,2010-11-11 Handbook of Modern Pharmaceutical Analysis Second Edition synthesizes the complex research and recent changes in the field while covering the techniques and technology required for today s laboratories The work integrates strategy case studies methodologies and implications of new regulatory structures providing complete coverage of quality assurance from the point of discovery to the point of use Treats pharmaceutical analysis PA as an integral partner to the drug development process rather than as a service to it Covers method development validation selection testing modeling and simulation studies combined with advanced exploration of assays impurity testing biomolecules and chiral separations Features detailed coverage of QA ethics and regulatory guidance quality by design good manufacturing practice as well as high tech methodologies and technologies from lab on a chip to LC MS LC NMR and LC NMR MS **Modern Inorganic Chemistry** William L. Jolly,1984

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