

KEY

35. Calculate the *maximum concentration* of fluoride ion possible in a solution in which $[Sr^{2+}] = 5.0 \times 10^{-3} M$.



$$K_{sp} = [Sr^{2+}][F^{-}]^2$$

$$[F^{-}] = \sqrt{\frac{K_{sp}}{[Sr^{2+}]}} = \sqrt{\frac{4.3 \times 10^{-9}}{5.0 \times 10^{-3}}} = 9.3 \times 10^{-4} M \quad \text{Answer} \quad [F^{-}] = 9.3 \times 10^{-4} M$$

36. Predict what would happen to the *solubility* of $PbSO_4(s)$ if some K_2SO_4 solution is added. Explain your answer and include the use of an equilibrium equation.



When K_2SO_4 (soluble) is added $[SO_4^{2-}]$ increases causing a shift to the LEFT which decreases the solubility of $PbSO_4$.

37. The solubility of $ZnCO_3$ in water is quite low. What could you add to increase the

solubility? Add an acid or Add a solution which precipitates Zn^{2+} or CO_3^{2-}

Explain fully how your method would work. Include the use of equilibrium equations.



H^+ ↓ Addition of an acid will supply H^+ , forming H_2O & CO_2 and decreasing the $[CO_3^{2-}]$. This causes a shift to the right which increases the solubility of $ZnCO_3$.

$$H_2CO_3 \rightarrow H_2O + CO_2(aq)$$


S^{2-} ↓ Adding Na_2S will supply S^{2-} which will precipitate the Zn^{2+} , decreasing $[Zn^{2+}]$ thus causing a shift to the RIGHT which increases the solubility of $ZnCO_3(s)$.

$$ZnS(s)$$

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Table of Contents Nelson Chemistry 12 Chapter 5 Solutions

1. Understanding the eBook Nelson Chemistry 12 Chapter 5 Solutions
 - The Rise of Digital Reading Nelson Chemistry 12 Chapter 5 Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry 12 Chapter 5 Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Chemistry 12 Chapter 5 Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry 12 Chapter 5 Solutions
 - Personalized Recommendations
 - Nelson Chemistry 12 Chapter 5 Solutions User Reviews and Ratings
 - Nelson Chemistry 12 Chapter 5 Solutions and Bestseller Lists
5. Accessing Nelson Chemistry 12 Chapter 5 Solutions Free and Paid eBooks
 - Nelson Chemistry 12 Chapter 5 Solutions Public Domain eBooks

- Nelson Chemistry 12 Chapter 5 Solutions eBook Subscription Services
- Nelson Chemistry 12 Chapter 5 Solutions Budget-Friendly Options
- 6. Navigating Nelson Chemistry 12 Chapter 5 Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Nelson Chemistry 12 Chapter 5 Solutions Compatibility with Devices
 - Nelson Chemistry 12 Chapter 5 Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nelson Chemistry 12 Chapter 5 Solutions
 - Highlighting and Note-Taking Nelson Chemistry 12 Chapter 5 Solutions
 - Interactive Elements Nelson Chemistry 12 Chapter 5 Solutions
- 8. Staying Engaged with Nelson Chemistry 12 Chapter 5 Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Chemistry 12 Chapter 5 Solutions
- 9. Balancing eBooks and Physical Books Nelson Chemistry 12 Chapter 5 Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Chemistry 12 Chapter 5 Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Nelson Chemistry 12 Chapter 5 Solutions
 - Setting Reading Goals Nelson Chemistry 12 Chapter 5 Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nelson Chemistry 12 Chapter 5 Solutions
 - Fact-Checking eBook Content of Nelson Chemistry 12 Chapter 5 Solutions
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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