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Question 6

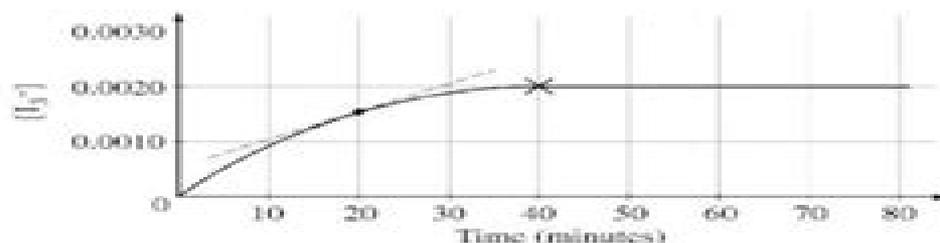
(10 points)

(a) The reaction is an oxidation-reduction (redox) reaction. 1 point

To go from I^- to I_3^- , the iodine has to be oxidized, so credit is earned for the formula of any reasonable oxidizing agent (e.g., F_2 , Co^{3+} , O_2 , MnO_4^-) 1 point

(b) (i) One point is earned for an "X" drawn on the curve at a point corresponding to approximately 37-42 minutes (see graph below). This is where the curve levels off. 1 point

(ii) The instantaneous rate of the reaction at 20 minutes is the slope of the line tangent to the curve at 20 minutes. (See the graph below.) 1 point



- No credit is earned if the student confuses "rate" with "rate constant".

(b) Keep the initial concentration of $S_2O_8^{2-}$ constant, change the initial concentration of I^- by a known amount, and see what effect this change has on the measured rate of the reaction. This procedure should be repeated, but this time keeping the concentration of I^- constant and changing the concentration of $S_2O_8^{2-}$. 2 points

NOTE: Points are earned for a method that involves controlling one reactant concentration and varying the other while comparing the reaction rates.

- No credit is earned if student confuses product and reactant concentrations.

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